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oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + \text{I}_2\text{O}_5$ $5 \text{ Co} + \text{I}_2 \rightarrow 5 \text{ Co}^{2+} + \text{I}_2\text{O}_5$

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reducing agent also 2sr o2 2sro sr0 to sr2 oxidized reducing agent o0 to o2 reduced ox ag 2li s li2s li0 to li1 oxidized red ag s0 to s2 reduced ox ag 3 2cs br2

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