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1 1 n2 3 f2 2 nf3 2 2 c6h10 17 o2 12
co2 10 h2o 3 4 5 1 hbr 1 khco3 1 h2o 1 kbr 1 co2 2 gabr3 3 na2so3 1 ga2 so3 3 6 nabr 3 sno 2 nf3 3 snf2 1 n2o3 using the
following equation 2 naoh h2so4 2 h2o na2so4

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assuming CeCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{Br}$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
br using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{O}_3 + \text{NaBr} + \text{SO}_2$

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 $\text{SiO}_2 + \text{C} \rightarrow \text{SiC} + \text{CO}$

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same thing goes with the O_3 i am so lost

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