



**GCE**

**Chemistry A**

Unit **F325**: Equilibria, Energetics and Elements

Advanced GCE

**Mark Scheme for June 2014**

# Ocr Chemistry A June 2014 F32paper

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$
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web start with 8 3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation 8 3 mole h<sub>2</sub>s x 2 mole h<sub>2</sub>o 2

mole  $\text{H}_2\text{S}$  x 18g  $\text{H}_2\text{O}$  i mole  $\text{H}_2\text{O}$  149 4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137}{149} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3}{32} = 0.666$  moles  $\text{O}_2$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al x  $\frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol} = 81.2 \text{ g}$  c determine the percent yield  $\frac{101.96}{160.0} \times 100 = 63.7\%$

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H<sub>2</sub> are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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