

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, Cl, O

3. Which atom has the higher electronegativity: sodium or sulfur?

Sulfur (sulfur is to the right of sodium and closer to fluorine)

4. Which two elements in the same family generally have similar properties?

There have been two cases here where a strong periodic trend is present.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest atomic radius? Which element has the smallest atomic radius? Which element has the greatest electronegativity?

Period 7, francium, fluorine, astatine, fluorine

7. Which trend in ionization energy values across a period is expected and the periodic table? Which element has the highest ionization energy? Which element has the lowest ionization energy?

Increases across the periodic table, fluorine has the highest, francium has the lowest.

8. Write the atoms in each group that have the largest atomic radius.

a. **Li**, **Na**, **K**
 b. **Be**, **Mg**, **Ca**
 c. **B**, **Al**, **Ga**
 d. **C**, **Si**, **Ge**
 e. **N**, **P**, **As**
 f. **O**, **S**, **Se**

g. **Br**, **I**, **At**
 h. **Cl**, **Br**, **I**
 i. **Li**, **K**, **Rb**
 j. **Na**, **K**, **Rb**

9. Write the atoms in each group that have the smallest atomic radius.

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