

# of substance:

**Substance:**  
 mass of particles of a substance  
 in mol

$6.02 \times 10^{23}$  particles  
 amount of substance

**Constant:**  
 particles per mol,  $6.02 \times 10^{23} \text{ mol}^{-1}$

**gas:**  
 mol measured in gas  
 equivalent to  $\text{M}^3$

**Molar Gas Volume:**  
 volume of gas per mole of gas molecules in a stated temperature + pressure  
 - Measured in  $\text{dm}^3 \text{ mol}^{-1}$  or  $\text{m}^3 \text{ mol}^{-1}$   
 At RTP (20°C and 100kPa):  
 (293K and 100000 Pa)  
 is  $24 \text{ dm}^3 \text{ mol}^{-1}$   
 at RTP.

Volume /  
 $\text{dm}^3$

**Electronegativity:** → the measure of attraction of a bonded atom in a covalent bond

**ELECTRONEGATIVITY INCREASES ACROSS PERIOD:**  
 - Increase in no. of protons in an atom  
 → Nuclear charge increases

**ELECTRONEGATIVITY DE:**  
 - More electron shells  
 → More shielding  
 → Atomic radius increases  
 → Nuclear charge less effective  
 → Nuclear attraction less  
 → Electronegativity

# OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS

## Atomic Structure

**Electrons:**  
 → Share p and s but different  
 shells different sizes

**Physical Properties:**  
 → Can be different - e.g. different  
 melting + boiling points  
 → i.e. massy water has higher melting  
 point + boiling point + density

**Chemical Properties:**  
 → Same no. of electrons + same electronic  
 structure in outer shell + transfer of  $e^-$  is  
 unaffected hence have same chemical  
 properties.  
 → Different no. of  $e^-$  has no effect  
 on bonding + reactions



→ Centre of a small dense  
 nucleus of protons + neutrons  
 surrounded by electrons orbiting  
 in electron orbitals in shells;  
 - No. of  $p^+$  = No. of  $e^-$   
 → Equal + opposite charges cancel  
 so atom has neutral charge  
 - Neutrons important in nuclear  
 stability - electrostatic repulsion  
 between  $p^+$  in small dense nucleus

**Mass Number:**  
 → No. of  $p^+$  + n in nucleus of atom  
 -  $p^+$  and  $n$  have relative mass of 1  
 and  $e^-$  mass is negligible to 0

**Relative Isotopic Mass:**  
 → The mass of an atom of an isotope  
 relative to  $1/12$ th the mass of an atom of  $C-12$

**Relative Atomic Mass:**  
 → The weighted mean mass of an atom of  
 an element relative to  $1/12$ th the mass of  
 an atom of  $C-12$

**Weighted Mean Mass:**  
 → Average mass of all isotopes, taking into  
 account the relative abundance of each isotope

- If dipole  
 moment is  
 out + molts

e.g.  $H_2O$

- O is more electronegative than H  
 - O-H has polar bond so not as  
 there is an over

e.g.  $CO_2$

O is more electronegative than C  
 C=O has a dipole  
 with bond

## Compounds + ions Formulae:

Alkali  $\text{NO}_3^-$  Halogenoalkane  $\text{RNO}_2^-$

# Ocr Chemistry Unit 5 June 2014

**Michal Rosen-Zvi**



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