

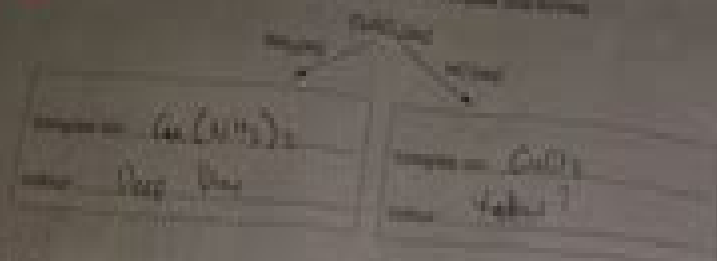
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a complex compound.

In the boxes, write the formula and name of the complex and ligand.



(b) Cu^{2+} ions form a complex ion B with two ethylenediamine ions and two water molecules. The ethylenediamine ion is a bidentate ligand.

The chemical formula of the ethylenediamine ion is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term bidentate ligand?

A ligand with two donor atoms forming two coordinate bonds with central metal ion.

(d) The complex ion A is shown in Fig. 1.2 below.

The structure of one of the ethylenediamine ions is shown below. The charge has been given.



Complex A

Complete the 3D diagram of the other two ethylenediamine ions of A. You do not need to include any charges.

Indicate with ticks whether the ethylenediamine ions are cis, trans, optical or a combination of these types.

Diagram														
Type	<table><tr><td>cis</td><td></td></tr><tr><td>trans</td><td>✓</td></tr><tr><td>optical</td><td></td></tr></table>	cis		trans	✓	optical		<table><tr><td>cis</td><td>✓</td></tr><tr><td>trans</td><td></td></tr><tr><td>optical</td><td>✓</td></tr></table>	cis	✓	trans		optical	✓
cis														
trans	✓													
optical														
cis	✓													
trans														
optical	✓													

(e) What is the empirical formula, including the charge, of the complex ion A?

$[Cu(NH_2)_2(en)_2]^{2+}$

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