

16. The dissociation of water is shown below:



- At 25 °C, the ion product of water, K_w , is 1.0×10^{-14} mol² dm⁻⁶.
At 65 °C, the ion product of water, K_w , is 1.0×10^{-13} mol² dm⁻⁶.

17. Explain what the dissociation of water is an indication of water's chemical nature.

Endothermic.

18. Predict, using a calculation, whether a solution of 1 mol dm⁻³ of acetic acid is acidic.

Acids

Explain by calculating pH.

19. pK_a , pH and pOH are significant notions.

Calculate pK_a of H_2O .

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-14}) = 14.00$$

$$pK_w = 14.00$$

20. A solution of hydrochloric acid is mixed with water and the resulting solution is diluted at 25 °C.

Calculate the pH of the resulting solution of water at 25 °C.

Give your answer to two decimal places.

$$pH = 1.30$$

pH = _____ (1)

(Total: 10)

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DJ Losen



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