

2.1.4 & 2.1.5 - Acids and redox

What are standard solutions?	A solution for which its concentration is known accurately.
What 3 things should you ensure when making standard solutions?	1. Ensure you have clean dry equipment. 2. Make washings to remove any salt BEFORE MAKING up the solution to the specified amount. 3. Invert the volumetric flask to ensure everything is dissolved.
How does ammonia react with acid?	• ammonia + acid $\rightarrow$ ammonium salt • $NH_3(aq) + HNO_3(aq) \rightarrow NH_4NO_3(aq)$
What are strong and weak acids ? Give 3 strong acids and 1 weak acid	• Strong acids dissociate fully to release H^+ ions in a solution e.g. $HCl \rightarrow H^+ + Cl^-$ • Weak acids dissociate partially to release H^+ ions in a solution – exist in equilibrium • Strong acids include: HCl, H_2SO_4, and HNO_3. • Weak acids include carboxylic acids such as CH_3COOH.
What are strong and weak bases? Give 1 weak base	• Strong bases dissociate fully to release OH^- ions in a solution • Weak bases dissociate partially to release OH^- ions in a solution – exist in equilibrium e.g. NH_3
What are diprotic and triprotic acids (with examples) and what does this mean for neutralisation?	• Diprotic acids and triprotic acids donate two and three protons respectively. • Examples include: H_2SO_4 and H_3PO_4. • They need to react to the number of moles of base to be neutralised

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Lei Huang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Lei Huang'.

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Table of Contents Ocr Mark Schemes Chemistry May 2013

1. Understanding the eBook Ocr Mark Schemes Chemistry May 2013
 - The Rise of Digital Reading Ocr Mark Schemes Chemistry May 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Mark Schemes Chemistry May 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Mark Schemes Chemistry May 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Mark Schemes Chemistry May 2013
 - Personalized Recommendations
 - Ocr Mark Schemes Chemistry May 2013 User Reviews and Ratings
 - Ocr Mark Schemes Chemistry May 2013 and Bestseller Lists

5. Accessing Ocr Mark Schemes Chemistry May 2013 Free and Paid eBooks
 - Ocr Mark Schemes Chemistry May 2013 Public Domain eBooks
 - Ocr Mark Schemes Chemistry May 2013 eBook Subscription Services
 - Ocr Mark Schemes Chemistry May 2013 Budget-Friendly Options
6. Navigating Ocr Mark Schemes Chemistry May 2013 eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Mark Schemes Chemistry May 2013 Compatibility with Devices
 - Ocr Mark Schemes Chemistry May 2013 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Mark Schemes Chemistry May 2013
 - Highlighting and Note-Taking Ocr Mark Schemes Chemistry May 2013
 - Interactive Elements Ocr Mark Schemes Chemistry May 2013
8. Staying Engaged with Ocr Mark Schemes Chemistry May 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Mark Schemes Chemistry May 2013
9. Balancing eBooks and Physical Books Ocr Mark Schemes Chemistry May 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Mark Schemes Chemistry May 2013
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Mark Schemes Chemistry May 2013
 - Setting Reading Goals Ocr Mark Schemes Chemistry May 2013
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Mark Schemes Chemistry May 2013
 - Fact-Checking eBook Content of Ocr Mark Schemes Chemistry May 2013
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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