

1 This question is about reversible reactions.

Sulfur trioxide is produced in a reversible reaction.

The equation for the reaction is:



1.1 What is the correct balanced equation for the reaction?

Tick **one** box.

- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
- ☐ $\text{SO}_2 + 2\text{O} \rightleftharpoons \text{SO}_3$
- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$
- ☐ $2\text{SO}_2 + 2\text{O}_2 \rightleftharpoons 2\text{SO}_3$

[1 mark]

1.2 The reaction of sulfur dioxide and oxygen to form sulfur trioxide is exothermic.

Which of the following statements is true?

Tick **one** box.

- ☐ The activation energy for the reaction of sulfur dioxide with oxygen is negative.
- ☐ Sulfur trioxide is colder than sulfur dioxide and oxygen.
- ☐ Heating the reaction will cause more sulfur trioxide to form.
- ☐ The reaction of sulfur trioxide to form sulfur dioxide and oxygen is endothermic.

[1 mark]

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Daniela Niemeyer



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Table of Contents Ocr Chemistry June 2013 Past Paper F3

1. Understanding the eBook Ocr Chemistry June 2013 Past Paper F3
 - The Rise of Digital Reading Ocr Chemistry June 2013 Past Paper F3
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry June 2013 Past Paper F3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry June 2013 Past Paper F3
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry June 2013 Past Paper F3
 - Personalized Recommendations
 - Ocr Chemistry June 2013 Past Paper F3 User Reviews and Ratings
 - Ocr Chemistry June 2013 Past Paper F3 and Bestseller Lists
5. Accessing Ocr Chemistry June 2013 Past Paper F3 Free and Paid eBooks
 - Ocr Chemistry June 2013 Past Paper F3 Public Domain eBooks
 - Ocr Chemistry June 2013 Past Paper F3 eBook Subscription Services
 - Ocr Chemistry June 2013 Past Paper F3 Budget-Friendly Options
6. Navigating Ocr Chemistry June 2013 Past Paper F3 eBook Formats

- ePub, PDF, MOBI, and More
 - Ocr Chemistry June 2013 Past Paper F3 Compatibility with Devices
 - Ocr Chemistry June 2013 Past Paper F3 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry June 2013 Past Paper F3
 - Highlighting and Note-Taking Ocr Chemistry June 2013 Past Paper F3
 - Interactive Elements Ocr Chemistry June 2013 Past Paper F3
 8. Staying Engaged with Ocr Chemistry June 2013 Past Paper F3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry June 2013 Past Paper F3
 9. Balancing eBooks and Physical Books Ocr Chemistry June 2013 Past Paper F3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry June 2013 Past Paper F3
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ocr Chemistry June 2013 Past Paper F3
 - Setting Reading Goals Ocr Chemistry June 2013 Past Paper F3
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ocr Chemistry June 2013 Past Paper F3
 - Fact-Checking eBook Content of Ocr Chemistry June 2013 Past Paper F3
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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