

16. The dissociation of water is shown below:



- At 25 °C, the ion product of water, K_w , is 1.0×10^{-14} .
At 65 °C, the ion product of water, K_w , is 1.0×10^{-13} .

Endothermic

17. Predict, using a calculation, whether a solution of 1.0 M HCl is neutral, acidic or basic. (3)

Acidic

Explain by calculating pH.

18. pK_a , pH and pOH are significant notions.

Calculate pK_a of HCl .

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-14}) = 13.02$$

$$pK_a = 13.02$$

19. A solution of hydrochloric acid is mixed with water and the resulting solution is neutral at 25 °C.

Calculate the pH of the resulting solution at 65 °C.

Give your answer to two decimal places.

$$pH = 13.02$$

pH = _____ (3)

(Total: 18)

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