

- (b) A solution is prepared by the addition of 10.0 mmol of CH_3COOH to 10.0 mmol of CH_3COONa and enough water to make 1.00 L of solution.

- (c) The solution reacts with a proper standard solution of NaOH to give a pH of 4.74.

How many moles of CH_3COOH are in the standard solution that reacts with NaOH to give a pH of 4.74?

In your answer, you should provide details of the chemical reaction that would be carried out, including appropriate equations and molar calculations.

Let the CH_3COOH be x moles and the CH_3COONa be y moles.

$$x + y = 10^{-3} + 10^{-3} = 2 \times 10^{-3} \text{ mol}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} = \frac{y}{x} \quad [\text{CH}_3\text{COOH}] = \frac{y}{K_a} = \frac{10^{-3}}{1.8 \times 10^{-5}} = 5.56 \times 10^{-2} \text{ mol/L}$$

$$\text{pH} = 4.74 \Rightarrow [\text{H}^+] = 1.82 \times 10^{-5} \text{ mol/L}$$

$$\text{pH} = 4.74 \Rightarrow \text{pH} = 4.74 \Rightarrow \text{pH} = 4.74 \Rightarrow \text{pH} = 4.74$$

Reaction: $\text{CH}_3\text{COOH} + \text{NaOH} \rightarrow \text{CH}_3\text{COONa} + \text{H}_2\text{O}$
 Initial: 10.0 mmol, 0.0 mmol, 10.0 mmol, 0.0 mmol
 Change: -x, +x, -x, +x
 Equilibrium: 10.0 - x, x, 10.0 - x, x
 pH = 4.74

- (b) The solution is prepared by the addition of 10.0 mmol of CH_3COOH to 10.0 mmol of CH_3COONa and enough water to make 1.00 L of solution.

How many moles of CH_3COOH are in the standard solution that reacts with NaOH to give a pH of 4.74?



- (c) The solution reacts with a proper standard solution of NaOH to give a pH of 4.74.

How many moles of CH_3COOH are in the standard solution that reacts with NaOH to give a pH of 4.74?

- (d) The solution reacts with a proper standard solution of NaOH to give a pH of 4.74.

$$\text{pH} = 4.74 \Rightarrow [\text{H}^+] = 1.82 \times 10^{-5} \text{ mol/L}$$

$$\text{pH} = 4.74 \Rightarrow \text{pH} = 4.74 \Rightarrow \text{pH} = 4.74 \Rightarrow \text{pH} = 4.74$$

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