

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions, as the reaction can proceed in either direction. Once the equilibrium is reached the forward and reverse reactions continue at equal rates and the concentrations of reactants and products remain constant. The importance of understanding that equilibrium is a dynamic state is that it allows us to predict the behavior of chemical systems. For example, if a reaction is at equilibrium, then the forward and reverse reactions continue to occur over time, but there is no net change in the concentrations of reactants and products. Equilibrium conditions are reached only at constant pressure, or concentration, or volume. Le Chatelier's Principle states that applying stress changes the equilibrium position of a system in order to minimize the effect of the change and bring the system back to equilibrium. When more products are added, an equilibrium reaction will shift to the left to consume the added products and bring the system back to equilibrium. When more reactants are added, the equilibrium will shift to the right to consume the added reactants and bring the system back to equilibrium. In this lab, you will be observing the changes in equilibrium when the amount of reactants, and temperature are changed. Follow Le Chatelier's Principle in the following reactions. Shifted equilibrium position denoted by the color change in solution according to the color indicators in the equations below. For example, if the color changes to red, the reaction shifts to the right to consume the added reactants and bring the system back to equilibrium.



Adding NaOH solution to $\text{Fe}^{3+}(\text{aq})$ and $\text{SCN}^{-}(\text{aq})$ will shift the equilibrium to the right to consume the added reactants and bring the system back to equilibrium. Adding the reactant increases the concentration of the reactant and bring the system back to equilibrium by producing more product and bring the system back to equilibrium.

Procedure (This Group):

- | | |
|---|---|
| 1. Weigh 0.5 g of FeCl_3 . | 1. Weigh 0.5 g of NaOH . |
| 2. Dissolve 0.5 g of FeCl_3 in 10 mL of water. | 2. Dissolve 0.5 g of NaOH in 10 mL of water. |
| 3. Add 10 mL of 0.1 M SCN^{-} to the FeCl_3 solution. | 3. Add 10 mL of 0.1 M SCN^{-} to the NaOH solution. |
| 4. Observe the color change. | 4. Observe the color change. |
| 5. Add 10 mL of 0.1 M SCN^{-} to the FeCl_3 solution. | 5. Add 10 mL of 0.1 M SCN^{-} to the NaOH solution. |
| 6. Observe the color change. | 6. Observe the color change. |
| 7. Add 10 mL of 0.1 M SCN^{-} to the FeCl_3 solution. | 7. Add 10 mL of 0.1 M SCN^{-} to the NaOH solution. |
| 8. Observe the color change. | 8. Observe the color change. |
| 9. Add 10 mL of 0.1 M SCN^{-} to the FeCl_3 solution. | 9. Add 10 mL of 0.1 M SCN^{-} to the NaOH solution. |
| 10. Observe the color change. | 10. Observe the color change. |

Observing Chemical Equilibrium Lab 46

Yicheng Fang



Observing Chemical Equilibrium Lab 46:

Unveiling the Magic of Words: A Report on "**Observing Chemical Equilibrium Lab 46**"

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