

CHEMISTRY Objective Final Exam Review 10-11

There are 50 questions on the test. The chapters are listed below. Each of the letters represents a concept/idea that you should know or look up.

1. CH 6: Periodic trends, where do they increase & decrease
 - a. Atomic radius & Ionic radius
 - b. Electronegativity
 - c. Electron affinity
 - d. Ionization energy
 - e. Reactivity of metals and of non-metals
2. CH 7: Look up properties of elements per group and the names of each group.
 - a. Group 1 very reactive in water; group 7 most reactive non-metals etc.
 - b. Fluorine most reactive in periodic table
 - c. Group 1 will lose e- to become +1
 - d. Elements in same group have same chem properties.
3. CH 8: Ionic compounds and bonds
 - a. Why do atoms bond? How many electrons make them stable?
 - b. How to name ionic compounds, go from name to formula
 - c. What elements make up ionic compounds?
 - d. Properties of ionic compounds
 - e. What is a crystal lattice and the arrangement.
 - f. What is metallic bonding?
4. CH 9: Covalent compounds
 - a. What elements make up covalent compounds?
 - b. Compare ionic vs covalent bonds. An e-neg difference of 1.7 is dividing factor.
 - c. What are pi bonds and sigma bonds?
 - d. Bond length Vs. Bond strength, and Bond dissociation energy.
 - e. What are molecules?
 - f. What are the common diatomic molecules?
 - g. Polar molecules, partial charges, what are they?
5. CH 10: Identify the types of chemical reactions
 - a. Be able to write the products for a double replacement and single replacement reactions.
 - b. Balancing equations
 - c. Writing equations from word names.
6. CH 11: The mole
 - a. Determine the molar mass of a compound
 - b. Calculate the amt of particles, ions, atoms, formula units, molecules in a compound.
 - c. Converting mole $\rightarrow$ mole, mole $\rightarrow$ grams, mole $\rightarrow$ particles and visa versa
 - d. Converting grams $\rightarrow$ mole, grams $\rightarrow$ grams, grams $\rightarrow$ particles visa vera
7. CH 12:
 - a. Stoichiometry conversions using balanced chemical equation
 - b. Mole $\rightarrow$ mole, mole $\rightarrow$ grams, grams $\rightarrow$ grams

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AN Whitehead



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Table of Contents Objective Final Review For Chemistry

1. Understanding the eBook Objective Final Review For Chemistry
 - The Rise of Digital Reading Objective Final Review For Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Objective Final Review For Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Objective Final Review For Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Objective Final Review For Chemistry
 - Personalized Recommendations
 - Objective Final Review For Chemistry User Reviews and Ratings
 - Objective Final Review For Chemistry and Bestseller Lists
5. Accessing Objective Final Review For Chemistry Free and Paid eBooks
 - Objective Final Review For Chemistry Public Domain eBooks
 - Objective Final Review For Chemistry eBook Subscription Services
 - Objective Final Review For Chemistry Budget-Friendly Options
6. Navigating Objective Final Review For Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Objective Final Review For Chemistry Compatibility with Devices
 - Objective Final Review For Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Objective Final Review For Chemistry
 - Highlighting and Note-Taking Objective Final Review For Chemistry
 - Interactive Elements Objective Final Review For Chemistry
8. Staying Engaged with Objective Final Review For Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Objective Final Review For Chemistry
- 9. Balancing eBooks and Physical Books Objective Final Review For Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Objective Final Review For Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Objective Final Review For Chemistry
 - Setting Reading Goals Objective Final Review For Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Objective Final Review For Chemistry
 - Fact-Checking eBook Content of Objective Final Review For Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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