# Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

#### Unit 1 - Matter & Measurement - Chapters 1 & 2

 A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

% error =  $\frac{experimentat-accepted}{accepted} \times 100$  % error =  $\frac{4.2-5.0}{5.0} \times 100 = -16\%$  error

#### Unit 2 - Periodic Table - Chapter 5

2.

| Element         | Symbol  | Atomic<br># | Protons | Neutrons | Electrons | Mass# | Charge |
|-----------------|---------|-------------|---------|----------|-----------|-------|--------|
| Fungsten<br>ion | 194 W+2 | 7-4         | 74      | 110      | 72        | 184   | +2     |

 Using the following atomic radius values, predict the atomic radius of potassium (K): Li = 145 pm, Na = 180 pm, K = 7, Rb = 235 pm K = 220 pm

#### Unit 3 - Reactions/Compounds - Chapters 7 & 8

- 4. Classify the compound as ionic or covalent. Name the compound
  - Fe<sub>2</sub>O<sub>3</sub> iron (III) oxide (ionic roman numerals for transition metals, change ending to -ide)
  - b. Cl<sub>2</sub>O<sub>8</sub> dichlorine octaoxide (covalent use prefixes, change ending to -ide)
- 5. Classify the compound as ionic or covalent. Write the formula
  - a. Trisulfur Dinitride S<sub>3</sub>N<sub>2</sub> (covalent)
  - b. Zinc Hydroxide Zn(OH)2 (ionic look up ion charges and criss-cross) Zn<sup>-2</sup> OH<sup>-3</sup>
  - c. Chromium (III) nitrate Cr(NO<sub>3</sub>)<sub>3</sub> (ionic criss-cross) Cr<sup>23</sup> NO<sub>3</sub><sup>-1</sup>
- 6. Write the equation, balance it, and classify the type of reaction:

Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate

> 1 FeSO<sub>4</sub> + 2 NaOH → 1 Fe(OH)<sub>2</sub> + 1 Na<sub>2</sub>SO<sub>4</sub> double displacement (two ionic compounds reacting)

7. Balance the equation and classify the type of reaction:

1 Cu + 2 AgNO<sub>3</sub> → 1 Cu(NO<sub>3</sub>)<sub>2</sub> + 2 Ag Single displacement (element and a compound reacting)

#### Unit 4 - Moles - Chapter 3

8. Calculate the molar mass of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>5</sub>.

Fe (2)(55.9 g/mol) = 111.8 g/mol

S (3)(32.1 g/mol) = 96.3 g/mol

+ O (12)(16.0 g/mol) = +192 g/mol

= 400.1 g/mol

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**JS Bruner** 

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#### **Table of Contents Review Sheet For Chemistry Final Answer Sheet**

- 1. Understanding the eBook Review Sheet For Chemistry Final Answer Sheet
  - The Rise of Digital Reading Review Sheet For Chemistry Final Answer Sheet
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Review Sheet For Chemistry Final Answer Sheet
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Review Sheet For Chemistry Final Answer Sheet
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Review Sheet For Chemistry Final Answer Sheet
  - Personalized Recommendations
  - Review Sheet For Chemistry Final Answer Sheet User Reviews and Ratings
  - Review Sheet For Chemistry Final Answer Sheet and Bestseller Lists

- 5. Accessing Review Sheet For Chemistry Final Answer Sheet Free and Paid eBooks
  - Review Sheet For Chemistry Final Answer Sheet Public Domain eBooks
  - Review Sheet For Chemistry Final Answer Sheet eBook Subscription Services
  - Review Sheet For Chemistry Final Answer Sheet Budget-Friendly Options
- 6. Navigating Review Sheet For Chemistry Final Answer Sheet eBook Formats
  - o ePub, PDF, MOBI, and More
  - Review Sheet For Chemistry Final Answer Sheet Compatibility with Devices
  - Review Sheet For Chemistry Final Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Review Sheet For Chemistry Final Answer Sheet
  - Highlighting and Note-Taking Review Sheet For Chemistry Final Answer Sheet
  - Interactive Elements Review Sheet For Chemistry Final Answer Sheet
- 8. Staying Engaged with Review Sheet For Chemistry Final Answer Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Review Sheet For Chemistry Final Answer Sheet
- 9. Balancing eBooks and Physical Books Review Sheet For Chemistry Final Answer Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Review Sheet For Chemistry Final Answer Sheet
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Review Sheet For Chemistry Final Answer Sheet
  - Setting Reading Goals Review Sheet For Chemistry Final Answer Sheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Sheet For Chemistry Final Answer Sheet
  - Fact-Checking eBook Content of Review Sheet For Chemistry Final Answer Sheet
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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