

Name _____ Class _____ Date _____

Assessment

Quiz

Section: Calculating Quantities in Reactions

In the space provided, write the letter of the term or phrase that best answers the question.

- _____ 1. The chemical equation $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$ can be correctly interpreted as
- 1 molecule of $\text{P}_4\text{O}_{10} + 6$ molecules of H_2O form 4 molecules of H_3PO_4 .
 - 1 mol $\text{P}_4\text{O}_{10} + 6$ mol H_2O reacts to form 4 mol H_3PO_4 .
 - 1 g $\text{P}_4\text{O}_{10} + 6$ g H_2O produces 4 g H_3PO_4 .
 - Both (a) and (b)
- _____ 2. Stoichiometry is the branch of chemistry that deals with the _____ between elements and compounds as reactants and products in a chemical reaction.
- bonding
 - energy transfers
 - numerical relationship
 - physical interactions
- _____ 3. How many mole ratios can be correctly obtained from the following chemical equation?
- $$\text{Na}_2\text{S} + \text{Cd}(\text{NO}_3)_2 \rightarrow 2\text{NaNO}_3 + \text{CdS}$$
- 1
 - 4
 - 12
 - 16
- _____ 4. Molar mass can be used to relate
- mass of reactant to moles of reactant.
 - moles of product to mass of product.
 - mass of product to moles of product.
 - All of the above
- _____ 5. In the chemical reaction described by the equation
- $$4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s),$$
- the mole ratio of iron(III) oxide to iron is
- 1:1.
 - 1:2.
 - 2:3.
 - 4:1.

Quiz Section Calculating Quantities In Reactions

R Barnett



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