

FINAL EXAMANITION IN PHYSICAL SCIENCE

I. Multiple Choices: For each of the following items, choose the BEST answer. On the answer sheet provided, shade the bubble that corresponds to the letter of your choice.

- The chemical formula of caffeine is $C_8H_{10}N_4O_2$. What is its molecular mass?
A. 165 g/mol B. 172 g/mol C. 194 g/mol D. 208 g/mol
- People usually use hydrogen peroxide (H_2O_2 ; molar mass = 34.02 g/mol) to clean their wounds. If Ricardo Delasay used 1.0 g of H_2O_2 to clean his gunshot wound, how many moles of H_2O_2 did he use?
A. 0.035 mol B. 0.030 mol C. 0.029 mol D. 0.0025 mol
- Methane (CH_4 ; molar mass = 16.04 g/mol) is one of the greenhouses gases. If 32.08 g of CH_4 was released in the atmosphere, how many molecules of CH_4 were added in the atmosphere?
A. 1.20×10^{26} B. 1.20×10^{23} C. 6.02×10^{26} D. 6.02×10^{23}

For numbers 4 and 5, consider the problem below:

- Iron (III) oxide reacts with carbon monoxide to form iron and carbon dioxide. Which is the balanced chemical equation that represents the reaction?
A. $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ C. $2Fe_2O_3 + 3CO \rightarrow 4Fe + 3CO_2$
B. $Fe_2O_3 + CO \rightarrow FeCO_2$ D. $Fe_2O_3 + 3CO \rightarrow Fe + 3CO_2$
- What mass of iron (III) oxide is needed to produce 5.6 kg of iron?
A. 4.0 kg B. 8.0 kg C. 12 kg D. 16 kg

For numbers 6-8, consider the problem below:

Ammonium carbonate decomposes to form ammonia gas, carbon dioxide and water. The reaction is represented by the equation below:



- How many moles of ammonia are formed for every mole of the reactant used?
A. 1 mole B. 2 moles C. 2.5 moles D. 3 moles
- How many moles of the reactant is required to produce 22 grams of CO_2 ?
A. 0.50 mole B. 1.0 mole C. 1.5 mole D. 2.0 moles
- How many grams of $(NH_4)_2CO_3$ is required to produce 1.5 moles of water?
A. 49 g B. 96 g C. 144 g D. 192 g

For numbers 9 and 10, consider the problem below:

45.9 g of K_2CO_3 was reacted with an excess of HCl , yielding 48.25 g of KCl .



- What is the theoretical yield?
A. 46.3 g B. 48.0 g C. 48.3 g D. 49.5 g
- What is the percent yield?
A. 92.5 % B. 93.5 % C. 97.0 % D. 97.5 %
- Which of the following will cause a soap to reduce its cleaning activity?
A. Alkyl hydrogen sulfate C. presence of surfactants
B. Hard water D. soft water
- Which of the following chemicals are linked to breast cancer in women?
A. Parabens B. PEG C. Paraffin wax D. BHA
- Which of the following disinfectants has a fungicidal property and is a potential carcinogen?

Physical Science Paper Final November 2014 Examanition Grade1

Amy J. Catalano, Michael A. Marino



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Table of Contents Physical Science Paper Final November 2014 Examanition Grade1

1. Understanding the eBook Physical Science Paper Final November 2014 Examanition Grade1
 - The Rise of Digital Reading Physical Science Paper Final November 2014 Examanition Grade1
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper Final November 2014 Examanition Grade1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Paper Final November 2014 Examanition Grade1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper Final November 2014 Examanition Grade1
 - Personalized Recommendations
 - Physical Science Paper Final November 2014 Examanition Grade1 User Reviews and Ratings
 - Physical Science Paper Final November 2014 Examanition Grade1 and Bestseller Lists

5. Accessing Physical Science Paper Final November 2014 Examanition Grade1 Free and Paid eBooks
 - Physical Science Paper Final November 2014 Examanition Grade1 Public Domain eBooks
 - Physical Science Paper Final November 2014 Examanition Grade1 eBook Subscription Services
 - Physical Science Paper Final November 2014 Examanition Grade1 Budget-Friendly Options
6. Navigating Physical Science Paper Final November 2014 Examanition Grade1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Paper Final November 2014 Examanition Grade1 Compatibility with Devices
 - Physical Science Paper Final November 2014 Examanition Grade1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Paper Final November 2014 Examanition Grade1
 - Highlighting and Note-Taking Physical Science Paper Final November 2014 Examanition Grade1
 - Interactive Elements Physical Science Paper Final November 2014 Examanition Grade1
8. Staying Engaged with Physical Science Paper Final November 2014 Examanition Grade1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Paper Final November 2014 Examanition Grade1
9. Balancing eBooks and Physical Books Physical Science Paper Final November 2014 Examanition Grade1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Paper Final November 2014 Examanition Grade1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Paper Final November 2014 Examanition Grade1
 - Setting Reading Goals Physical Science Paper Final November 2014 Examanition Grade1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Paper Final November 2014 Examanition Grade1
 - Fact-Checking eBook Content of Physical Science Paper Final November 2014 Examanition Grade1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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