

HALF-LIFE PROBLEMS

Name	Block
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1. An isotope of cesium (cesium-137) has a half-life of 30 years. If 1.0 g of cesium-137 disintegrates over a period of 90 years, how many g of cesium-137 would remain?

$$\frac{90}{30} = 3 \frac{1}{2}$$

$\begin{pmatrix} 1 \\ 0.5 \\ 2.5 \\ 12.5 \end{pmatrix} g$

2. Actinium-226 has a half-life of 29 hours. If 100 mg of actinium-226 disintegrates over a period of 58 hours, how many mg of actinium-226 will remain?

58-2.4.4

100
50
25

3. Sodium-25 was to be used in an experiment, but it took 3.0 minutes to get the sodium from the reactor to the laboratory. If 5.0 mg of sodium-25 was removed from the reactor, how many mg of sodium-25 were placed in the reaction vessel 3.0 minutes later if the half-life of sodium-25 is 60 seconds?

$$\frac{3.0}{1.0} = 3 \text{ k.i.}$$

$\left\{ \begin{array}{l} 5 \text{ mg} \\ 2.5 \\ 1.25 \\ 0.625 \text{ mg} \end{array} \right.$

4. The half-life of isotope X is 2.0 years. How many years would it take for a 4.0 mg sample of X to decay and have only 0.50 mg of it remain?

$$3 \times 2.0 \text{ yr} = 6 \text{ yr}$$

4.00
2.00
1.00
5.00 } 3 h-L

5. Selenium-83 has a half-life of 25.0 minutes. How many minutes would it take for a 10.0 mg sample to decay and have only 1.25 mg of it remain?

$$3 \times 25 = 75 \text{ mi}$$

1.25 3 h.L

6. The half-life of Po-218 is three minutes. How much of a 2.0 gram sample remains after 15 minutes? Suppose you wanted to buy some of this isotope, and it required half an hour for it reach you. How much should you order if you need to use 0.10 gram of this material?

$\frac{1}{2} \text{ mol}$

2.3
1.5
2.2
1.25
1.0625

30m, 10 h.f.

10 k.f

$.1 \times 2 \times 2 \times 2 \times 2 = 2 \times 2 \times 2 = 2 \times 2 \times 2$

$102.4g$

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Table of Contents Physical Science If8767 Half Life Calculations Answers

1. Understanding the eBook Physical Science If8767 Half Life Calculations Answers
 - The Rise of Digital Reading Physical Science If8767 Half Life Calculations Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Half Life Calculations Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science If8767 Half Life Calculations Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Half Life Calculations Answers
 - Personalized Recommendations
 - Physical Science If8767 Half Life Calculations Answers User Reviews and Ratings
 - Physical Science If8767 Half Life Calculations Answers and Bestseller Lists
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 - Adjustable Fonts and Text Sizes of Physical Science If8767 Half Life Calculations Answers
 - Highlighting and Note-Taking Physical Science If8767 Half Life Calculations Answers
 - Interactive Elements Physical Science If8767 Half Life Calculations Answers
- 8. Staying Engaged with Physical Science If8767 Half Life Calculations Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science If8767 Half Life Calculations Answers
- 9. Balancing eBooks and Physical Books Physical Science If8767 Half Life Calculations Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science If8767 Half Life Calculations Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science If8767 Half Life Calculations Answers
 - Setting Reading Goals Physical Science If8767 Half Life Calculations Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science If8767 Half Life Calculations Answers
 - Fact-Checking eBook Content of Physical Science If8767 Half Life Calculations Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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