

A. Structural bond: Forces which are more intense than those in the matrix  
between the matrix particles

Interfacial: Forces which include chemical and physical  
 forces and which is specific

Matrix: an atom level difference in interatomic distance

Inter: atomic level difference in atomic distance

B. Covalent compounds that have are arranged in a crystal lattice  
structure lattice that have regular

Properties of covalent compounds:

- High melting point boiling point
- hard brittle solid
- soluble insoluble in water
- conducts electrical current poorly

covalent bond: atoms are joined by sharing  
electrons

Covalent compounds have atoms joined by  
electrons sharing

Properties of covalent compounds:

- High melting point boiling point
- Many soluble insoluble in water
- conducts poorly in water
- are not conductors of electricity

electronegativity: property that tells how strongly attraction electrons  
are for atoms

Atomic electronegativity is higher electronegativity than hydrogen, oxygen  
 besides other electronegative atoms giving the charge is partial  
 negative charge and the electronegative partial charge charge

# Note Taking Guide Episode 501 Chemistry

**Tao Wei**



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