

Chemical Formula

Talking about proportions, molecules and relative of atoms present in a compound.

Molecular $\text{C}_{10}\text{H}_{22}\text{O}_4$ (molecular)

Formula represents the number of atoms of each element in a single molecule.

Empirical $\text{C}_5\text{H}_{11}\text{O}_2$ (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound.

Atomic Mass

The average mass of all atoms of an element is called its atomic mass.

Isotopes

^{12}C and ^{13}C are isotopes of carbon.

Isotopes have same chemical properties but different physical properties.

Radioisotopes

^{14}C is a radioisotope of carbon.

Radioisotopes have same chemical properties but different physical properties.

^{14}C is a radioisotope of carbon.

Note Taking Guide Episode 601 Chemical Formula Key

LL Leslie

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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