

### Chemical Formula

Talking about proportions, molecules and relative all atoms contained in a molecule.

Molecular  $\text{C}_{10}\text{H}_{18}\text{O}_2$  (molecular)

Formula represents the number of atoms of each element in a single molecule.

Empirical  $\text{C}_{10}\text{H}_{18}\text{O}_2$  (molecular)

Formula represents the proportions of atoms of positive and negative ions in a compound.

### Atomic Mass

The proportions of atoms in a compound are proportional to their relative atomic masses.

### Relative Atomic Mass

The relative atomic mass of an element is the average mass of an atom of that element.

The relative atomic mass of an element is the average mass of an atom of that element.

### Relative Molecular Mass

The relative molecular mass of a compound is the average mass of a molecule of that compound.

The relative molecular mass of a compound is the average mass of a molecule of that compound.

The relative molecular mass of a compound is the average mass of a molecule of that compound.

# Note Taking Guide Episode 601 Answers Physics

**RD Boyd**



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