

Worksheet #1: Significant Figures & Dimensional Analysis

Directions: For each problem below, write the equation and show your work. Be sure to **box your final answer**.

Part 1: Solve the following problems using scientific notation and rounding to the appropriate value.

- a. $300,000,000 \rightarrow 3.0 \times 10^8$
- b. $450,000 \rightarrow 4.5 \times 10^5$
- c. $0.000000000 \rightarrow 9.0 \times 10^{-9}$
- d. $-0.0000000 \rightarrow -3.0 \times 10^{-6}$

Part 2: Solve the following problems, and show your final answer with the appropriate number of significant figures.

- a. $1.2000 + 71.80 \rightarrow 73.0000 \rightarrow 73.00$
- b. $9.1 \times 10^3 + 20 \rightarrow 9100 + 20 \rightarrow 9120$
- c. $0.000000000 \rightarrow 0.000000000$
- d. $0.000000000 \rightarrow 0.000000000$
- e. $0.000000000 \rightarrow 0.000000000$
- f. $0.000000000 \rightarrow 0.000000000$
- g. $0.000000000 \rightarrow 0.000000000$
- h. $0.000000000 \rightarrow 0.000000000$
- i. $0.000000000 \rightarrow 0.000000000$
- j. $0.000000000 \rightarrow 0.000000000$
- k. $0.000000000 \rightarrow 0.000000000$
- l. $0.000000000 \rightarrow 0.000000000$
- m. $0.000000000 \rightarrow 0.000000000$
- n. $0.000000000 \rightarrow 0.000000000$
- o. $0.000000000 \rightarrow 0.000000000$
- p. $0.000000000 \rightarrow 0.000000000$
- q. $0.000000000 \rightarrow 0.000000000$
- r. $0.000000000 \rightarrow 0.000000000$
- s. $0.000000000 \rightarrow 0.000000000$
- t. $0.000000000 \rightarrow 0.000000000$
- u. $0.000000000 \rightarrow 0.000000000$
- v. $0.000000000 \rightarrow 0.000000000$
- w. $0.000000000 \rightarrow 0.000000000$
- x. $0.000000000 \rightarrow 0.000000000$
- y. $0.000000000 \rightarrow 0.000000000$
- z. $0.000000000 \rightarrow 0.000000000$

Part 3: For each problem below, show your work. Always use units and box in your final answer.

- a. The density of pure silver is 10.5 g/cm^3 at 20°C . If 5.25 g of pure silver pellets are added to a graduated cylinder containing 11.2 mL of water, to what volume level will the water in the cylinder rise?

$d = \frac{m}{V}$

$10.5 \text{ g/cm}^3 \times 1 \text{ cm}^3 = 10.5 \text{ g}$

$5.25 \text{ g} \div 10.5 \text{ g/cm}^3 = 0.5 \text{ cm}^3$

$11.2 \text{ mL} + 0.5 \text{ mL} = 11.7 \text{ mL}$

- b. The density of air at ordinary atmospheric pressure and 20°C is 1.29 g/L . What is the mass, in kilograms, of the air in a room that measures $12.5 \text{ m} \times 15.5 \text{ m} \times 8.80 \text{ m}$?

$d = 1.29 \text{ g/L}$

$V = 12.5 \times 15.5 \times 8.80$

$V = 1690 \text{ L}$

$m = ? \text{ kilograms}$

$1.29 \text{ g/L} \times 1690 \text{ L} = 2170 \text{ g}$

$2170 \text{ g} \div 1000 = 2.17 \text{ kg}$

Density is measured in kg/m^3 , and the volume is given in L . So you must convert $\text{L} \rightarrow \text{m}^3$. $1 \text{ L} = 0.001 \text{ m}^3$

$1 \text{ L} = 0.001 \text{ m}^3$

$1690 \text{ L} \times 0.001 = 1.69 \text{ m}^3$

$1.29 \text{ kg/m}^3 \times 1.69 \text{ m}^3 = 2.17 \text{ kg}$

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M Tight



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Table of Contents Northridge Learning Center Packet Answers Chemistry 2

1. Understanding the eBook Northridge Learning Center Packet Answers Chemistry 2
 - The Rise of Digital Reading Northridge Learning Center Packet Answers Chemistry 2
 - Advantages of eBooks Over Traditional Books
2. Identifying Northridge Learning Center Packet Answers Chemistry 2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Northridge Learning Center Packet Answers Chemistry 2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Northridge Learning Center Packet Answers Chemistry 2
 - Personalized Recommendations
 - Northridge Learning Center Packet Answers Chemistry 2 User Reviews and Ratings
 - Northridge Learning Center Packet Answers Chemistry 2 and Bestseller Lists
5. Accessing Northridge Learning Center Packet Answers Chemistry 2 Free and Paid eBooks
 - Northridge Learning Center Packet Answers Chemistry 2 Public Domain eBooks
 - Northridge Learning Center Packet Answers Chemistry 2 eBook Subscription Services
 - Northridge Learning Center Packet Answers Chemistry 2 Budget-Friendly Options
6. Navigating Northridge Learning Center Packet Answers Chemistry 2 eBook Formats

- ePub, PDF, MOBI, and More
 - Northridge Learning Center Packet Answers Chemistry 2 Compatibility with Devices
 - Northridge Learning Center Packet Answers Chemistry 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Northridge Learning Center Packet Answers Chemistry 2
 - Highlighting and Note-Taking Northridge Learning Center Packet Answers Chemistry 2
 - Interactive Elements Northridge Learning Center Packet Answers Chemistry 2
 8. Staying Engaged with Northridge Learning Center Packet Answers Chemistry 2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Northridge Learning Center Packet Answers Chemistry 2
 9. Balancing eBooks and Physical Books Northridge Learning Center Packet Answers Chemistry 2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Northridge Learning Center Packet Answers Chemistry 2
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Northridge Learning Center Packet Answers Chemistry 2
 - Setting Reading Goals Northridge Learning Center Packet Answers Chemistry 2
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Northridge Learning Center Packet Answers Chemistry 2
 - Fact-Checking eBook Content of Northridge Learning Center Packet Answers Chemistry 2
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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