# Stoichiometry Mixed Review

1. 
$$N_2 + 3H_2 \rightarrow 2NH_3$$

$$1 \text{ mol N}_2 = 2 \text{ mol NH}_3$$

$$25 g N_2 \times \frac{1 mol N_2}{28.0 g N_2} \times \frac{2 mol NH_3}{1 mol N_2} \times \frac{22.4 LNH_3}{1 mol NH_3} = 40 LNH_3$$

2. 
$$2KCIO_3 \rightarrow 2KCI + 3O_2$$

2 mol KClO<sub>3</sub> = 3 mol O<sub>2</sub>

$$5.0 g KClO_3 \times \frac{1 mol KClO_3}{122.6 g KClO_3} \times \frac{3 mol O_2}{2 mol KClO_3} \times \frac{32.0 g O_2}{1 mol O_2} = 2.0 g O_2$$

# **Review Stoichiometry Mixed Review**

**Vincent Dusastre** 

## **Review Stoichiometry Mixed Review:**

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