

CHEM 101 Quiz #6 Chapter 6 ANSWERS**1. (4 points)**

The specific heat (s) of iron (Fe) is 0.40 J/(g K) . Calculate the heat effect q when 20 g of iron is heated from 100°C to 200°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } ^\circ\text{C} \quad q = 20 \times 0.40 \times (200 - 100) = +800 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Mg = 24.0 g/mol .

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

$$2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol}) \times (-600)(\text{kJ/mol}) = -60 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following four reactions is a *formation reaction*? Circle the correct answer.



Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur so e.g. Cl_2 not Cl , and H_2 not H !

CHEM 101 Section L52 Quiz #6**1. (4 points)**

The specific heat (s) of copper (Cu) is 0.38 J/(g K) . Calculate the heat effect q when 10 g of copper is heated from 100°C to 300°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } ^\circ\text{C} \quad q = 10 \times 0.38 \times (300 - 100) = +760 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Ca = 40.0 g/mol .

Calculate the enthalpy change ΔH when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

$$8.0 \text{ g Ca} = 8.0/40 = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol}) \times (-650)(\text{kJ/mol}) = -130 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following reactions is a *formation reaction*? Circle the correct answer.



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S Ashworth



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practitioners administrators policymakers and advocates how and why this particular reform of conditions achieved successful outcomes when others failed Maintaining that the Chicago Crime Lab findings are the gold standard evidence based research EBR in pretrial detention Roush holds that the observed firsts for juvenile detention may perhaps have the power to transform all custody practices He shows that the findings validate a new model of institutional reform based on cognitive behavioral programming CBT reveal statistically significant reductions in in custody violence and recidivism and demonstrate that at least one variation of short term secure custody can influence positively certain life outcomes for Chicago s highest risk and most disadvantaged youth With the Quarterly Journal of Economics imprimatur and endorsement by the President s Council of Economic Advisors the book is a reverse engineering of these once in a lifetime events recidivism reduction and EBR in pretrial detention that explains the important and transformative implications for the future of juvenile justice practice The book is essential reading for graduate students in juvenile justice criminology and corrections as well as practitioners judges and policymakers

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