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Solutions Manual

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Chapter 7

Equilibria for Multiple Component Systems

- 7.2. With only a single component and phase, if only $\text{Fe}_2(\text{SO}_4)_3$ were present in our system, we would have $C - P + 2 = 1 - 1 + 2 = 2$ degrees of freedom.
- 7.4. For a single-component system, there will never be negative degrees of freedom. The maximum number of phases that can exist in equilibrium is three, so the Gibbs phase rule would yield $1 - 3 + 2 = 0$ degrees of freedom (like might occur at the triple point). For a two-component system, the maximum number of phases that can exist in equilibrium is four. There will never be negative degrees of freedom.
- 7.6. In this reaction, there are only three independent components (the amount of the fourth can be determined from the equilibrium constant) and 3 phases, two solids and a gas mixture. Therefore, there are $3 - 3 + 2 = 2$ degrees of freedom.
- 7.8. For water: $n = \frac{pV}{RT} = \frac{(23.76 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 6.39 \times 10^{-3} \text{ mol}$ of water is needed to ensure that there is a liquid and gas phase. This equals 0.115 grams.
For methanol:
- $$n = \frac{pV}{RT} = \frac{(125.0 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 3.36 \times 10^{-2} \text{ mol}$$
- of methanol is needed to ensure that there is a liquid and gas phase. This equals 1.08 grams.
- 7.10. Since, according to equation 7.11, $a_i = \frac{p_i}{p_i^*}$, and since $p_i^* = 760 \text{ torr}$, we can calculate a_i :

Physical Chemistry Ball Manual

Judd E. Hollander



Physical Chemistry Ball Manual:

Student Solutions Manual for Physical Chemistry David Ball, 2002-09 The manual consists of complete solutions to all odd end of chapter exercises and problems *Student Solutions Manual, Physical Chemistry Second Edition, David W. Ball* Jörg C. Woehl, 2015 **Physical Chemistry + Student Solutions Manual**, *Student Solutions Manual for Physical Chemistry* C. A. Trapp, Peter Atkins, Julio de Paula, 2009-12-18 With its modern emphasis on the molecular view of physical chemistry its wealth of contemporary applications vivid full color presentation and dynamic new media tools the thoroughly revised new edition is again the most modern most effective full length textbook available for the physical chemistry classroom Available in Split Volumes For maximum flexibility in your physical chemistry course this text is now offered as a traditional text or in two volumes Volume 1 Thermodynamics and Kinetics ISBN 1 4292 3127 0 Volume 2 Quantum Chemistry Spectroscopy and Statistical Thermodynamics ISBN 1 4292 3126 2 The Elements of Physical Chemistry John Livingston Rutgers Morgan, 1908 *Physical Chemistry for Electrical Engineers* John Livingston Rutgers Morgan, 1909 Soil Physical Chemistry Donald L. Sparks, 2018-02-06 *Soil Physical Chemistry Second Edition* takes up where the last edition left off With comprehensive and contemporary discussions on equilibrium and kinetic aspects of major soil chemical process and reactions this excellent text reference presents new chapters on precipitation dissolution modeling of adsorption reactions at the mineral water interface and the chemistry of humic substances An emphasis is placed on understanding soil chemical reactions from a microscopic point of view and rigorous theoretical developments such as the use of modern in situ surface chemical probes such as x ray adsorption fine structure XAFS Fourier transform infrared FTIR spectroscopies and scanning probe microscopies SPM are discussed *Student Solutions Manual for Ball's Physical Chemistry, 2nd* David W. Ball, 2014 Master problem solving using the detailed solutions in this manual which contains completely worked out solutions to all odd end of chapter exercises and problems **Handbook of Problems in Direct Fire** James Monroe Ingalls, 1903 A Handbook of ophthalmic science and practice Henry Edward Juler, 1884 *Atkins' Physical Chemistry* Peter William Atkins, Julio De Paula, James Keeler, 2023 Revised edition of Atkins Physical chemistry Peter Atkins Julio de Paula James Keeler Eleventh edition 2018 *The Practitioner's Handbook of Treatment* J. Milner Fothergill, 1887 **Physical and natural therapeutics** Georges Hayem, 1895 **Physical Chemistry** David Ball, 2002-08-20 Intended for the year long calculus based physical chemistry course for science and engineering majors PHYSICAL CHEMISTRY follows a traditional organization while concentrating on core topics The text does not cover some higher level topics for example photochemistry molecular beams thermal physics and polymers found in some textbooks and rarely covered in the undergraduate physical chemistry course but more fully explains the essential elements of the discipline Written by a dedicated chemical educator and researcher this text is intended for those students who are trying to learn physical chemistry a book that works as a textbook and not as an encyclopedia Where appropriate there is some focus on mathematical manipulations providing

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