

QUESTION 3

STEAM GENERATION P.M N5

Use a steam table or relevant formula for the following:

- 3.1 Feedwater at a flow rate of 1,2 kilogram per second is supplied at 27 °C. The specific heat capacity of water is 4,187 kJ/kg.K.

Calculate the enthalpy of the fluid at MJ/hr.

- 3.2 A plant operator is monitoring the enthalpy of wet steam leaving a boiler at 1,5 MPa. The steam flows at 0,9 kilograms per minute with a 3% moisture content.

NOVEMBER 2020 REVISION

What is the expected enthalpy of the steam after 3 minutes?

- 3.3 Determine the enthalpy of dry steam in kJ/kg at a pressure of 1200 kPa.
- 3.4 2 kg feed water enters a plant at a temperature of kJ/kg 31 °C and exits as a superheated steam delivered at 2,5 MPa and a temperature of 288 °C. The specific heat capacity of the superheated steam is 2,1 kJ/kg.K.

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper's mass into account add these to the beaker 3 measure 50 0 ml of water in the

cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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