

A. Identify an equilibrium, the K_{eq} expression for that equilibrium, and the K_{eq} as applied to a particular equilibrium.



Classroom: *As a class, we assume where the reaction is not at equilibrium, then you will not have equilibrium that is given. Therefore, it is important.*

A. Equilibrium:

• If the K_{eq} is > 1 , then the reaction is favored to the right (products are favored).

• If the concentration of products is greater than the concentration of reactants, then the reaction is favored to the right.



$$K_{eq} = \frac{[C][D]}{[A][B]}$$

Ex: K_{eq} is _____ reaction is favored

$K_{eq} > 1$	Products
$K_{eq} < 1$	Reactants
$K_{eq} = 1$	Equilibrium

Ex: 10% Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.



$$\frac{[H_2][CO_2]}{[H_2O][CO]}$$

$$\frac{[0.0001][0.0001]}{[0.0001][0.0001]}$$

$$K_{eq} = 0.0001$$

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Guillaume Favre



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The Wireless World ,1933 *New York Times Saturday Book Review Supplement* ,1993-04

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