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web for practical purposes a rounded version of the constant can be used in exams $6.02 \times 10^{23} \text{ mol}^{-1}$ the reason we have mol^{-1} as the unit is because this is the number of entities per mole of substance for example one mole of sodium na contains 6.02×10^{23} atoms of sodium one mole of hydrogen h₂ contains 6.02×10^{23} molecules of hydrogen