

Periods and Groups: Valence Electrons: Activity

Periods and Groups

Remember that both the outer number of valence electrons being smaller (periodic) or larger (transition groups) than the periodic table groups is indicated that both groups. They will not identify groups 1 & 2 & 10 & 11 & 12 as the outer period is shown.

Identify the group number in the table and write the number in the box provided in Periods, Groups, and Valence Electrons in the table provided.

		Groups														
		1A	2A	Transition Metals						3A	4A	5A	6A	7A	8A	
Period 1		1														
Period 2		2	3													
Period 3		3	4													
Period 4		4	5													

Looking at the table "identify" periods that are decreasing in period number.

What element is in Group 1, and Group 8A? Group 1 element is always that is element 119. The element is hydrogen.

What element is in Group 11A and Period 2?

When group number and is 4/10/11/12?

Group: _____ Period: _____

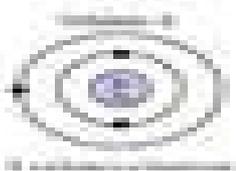
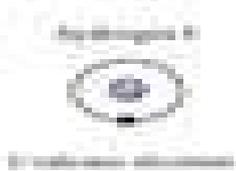
What element is in Group 2A and Period 3?

What element is hydrogen?

Group: _____ Period: _____

Valence Electrons

Valence Electrons are the outermost electrons in an atom. Each period contains the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.



Group 1, 2, 10, 11, 12 and 13 valence electrons. How many valence electrons will they have?

How many electrons?

Hydrogen and Helium are both in Group 1 & 2 respectively. How many electrons?

Group 13: Group 13 elements have 3 valence electrons. Elements in Group 13 and 14 are similar to Group 1 & 2. Group 13 elements have 3 valence electrons. "B, Al, Ga, In, Tl" Group 14 elements have 4 valence electrons. "C, Si, Ge, Sn, Pb" Group 15 elements have 5 valence electrons. "N, P, As, Sb, Bi" Group 16 elements have 6 valence electrons. "O, S, Se, Te, Po" Group 17 elements have 7 valence electrons. "F, Cl, Br, I, At" Group 18 elements have 8 valence electrons. "He, Ne, Ar, Kr, Xe, Rn"

Periods And Groups Valence Electrons Masses Answers

SJ Ball



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