

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

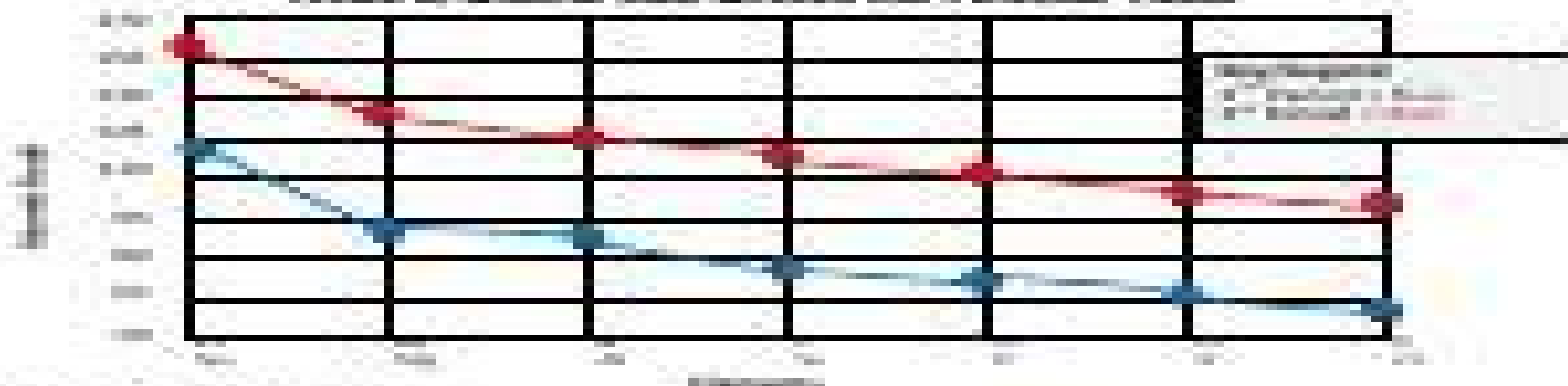
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, use the other graphs
3. Label the box as 'not suitable' if not suitable
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
2	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
3	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
4	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
5	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
6	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
7	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
8	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
9	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
10	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius

Now label the trends in the following table:

Trend in Atomic Size across the Periodic Table



Electronegativity Questions

1. Following the trend in atomic size across the periodic table, electronegativity
2. This means that atoms that are more electronegative (higher) have a stronger pulling force on the electrons in the atom.

Atomic Size Questions

1. Electronegativity increases from left to right, e.g. Na to Cl. This is the opposite movement to what you found in relation to the periodic table? Yes
2. This means that atoms that are more electronegative (higher) have a stronger pulling force on the electrons in the atom.

Periodicity Activity Okstate Answer Key

Lingsheng Yao

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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