

## Periodic Table of the Elements

| Periodic Table of the Elements |                |                |                |                |                |                |                |                |                |                |                |                |                |                |              |              |              |              |             |             |    |    |    |    |    |    |    |    |  |
|--------------------------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|--------------|--------------|--------------|--------------|-------------|-------------|----|----|----|----|----|----|----|----|--|
|                                | 1              | 2              |                |                |                |                |                |                |                |                |                |                | 3              | 4              | 5            | 6            | 7            | 8            | 9           | 10          | 11 | 12 | 13 | 14 | 15 | 16 | 17 | 18 |  |
| 1                              | H<br>1.008     | He<br>4.003    |                |                |                |                |                |                |                |                |                |                | Li<br>6.941    | Be<br>9.012    | B<br>10.81   | C<br>12.01   | N<br>14.01   | O<br>16.00   | F<br>18.998 | Ne<br>20.18 |    |    |    |    |    |    |    |    |  |
| 2                              | Na<br>22.99    | Mg<br>24.31    |                |                |                |                |                |                |                |                |                |                | Al<br>26.98    | Si<br>28.09    | P<br>30.97   | S<br>32.07   | Cl<br>35.45  | Ar<br>39.95  |             |             |    |    |    |    |    |    |    |    |  |
| 3                              | K<br>39.10     | Ca<br>40.08    | Sc<br>44.96    | Ti<br>47.88    | V<br>50.94     | Cr<br>52.00    | Mn<br>54.94    | Fe<br>55.85    | Co<br>58.93    | Ni<br>58.69    | Cu<br>63.55    | Zn<br>65.38    | Ga<br>69.72    | Ge<br>72.64    | As<br>74.92  | Se<br>78.96  | Br<br>79.90  | Kr<br>83.80  |             |             |    |    |    |    |    |    |    |    |  |
| 4                              | Rb<br>85.47    | Sr<br>87.62    | Y<br>88.91     | Zr<br>91.22    | Nb<br>92.91    | Mo<br>95.94    | Tc<br>98.91    | Ru<br>101.07   | Rh<br>102.91   | Pd<br>106.42   | Ag<br>107.87   | Cd<br>112.41   | In<br>114.82   | Sn<br>118.71   | Sb<br>121.76 | Te<br>127.60 | I<br>126.91  | Xe<br>131.29 |             |             |    |    |    |    |    |    |    |    |  |
| 5                              | Cs<br>132.91   | Ba<br>137.33   | La<br>138.91   | Ce<br>140.12   | Pr<br>140.91   | Nd<br>144.24   | Pm<br>144.91   | Sm<br>150.36   | Eu<br>151.96   | Gd<br>157.25   | Tb<br>158.93   | Dy<br>162.50   | Ho<br>164.93   | Er<br>167.26   | Tm<br>168.93 | Yb<br>173.05 | Lu<br>174.97 |              |             |             |    |    |    |    |    |    |    |    |  |
| 6                              | Fr<br>[223]    | Ra<br>[226]    | Ac<br>[227]    | Th<br>[232]    | Pa<br>[231]    | U<br>[238]     | Np<br>[237]    | Pu<br>[244]    | Am<br>[243]    | Cm<br>[247]    | Bk<br>[247]    | Cf<br>[251]    | Es<br>[252]    | Fm<br>[257]    | Mn<br>[288]  | Ni<br>[289]  | Pu<br>[289]  | At<br>[210]  | Fr<br>[223] |             |    |    |    |    |    |    |    |    |  |
| 7                              | Uu<br>[289]    | Uu<br>[291]    | Uu<br>[293]    | Uu<br>[295]    | Uu<br>[297]    | Uu<br>[299]    | Uu<br>[301]    | Uu<br>[303]    | Uu<br>[305]    | Uu<br>[307]    | Uu<br>[309]    | Uu<br>[311]    | Uu<br>[313]    | Uu<br>[315]    | Uu<br>[317]  | Uu<br>[319]  | Uu<br>[321]  | Uu<br>[323]  |             |             |    |    |    |    |    |    |    |    |  |
| 8                              | Ce<br>[140.12] | Pr<br>[140.91] | Nd<br>[144.24] | Pm<br>[144.91] | Sm<br>[150.36] | Eu<br>[151.96] | Gd<br>[157.25] | Tb<br>[158.93] | Dy<br>[162.50] | Ho<br>[164.93] | Er<br>[167.26] | Tm<br>[168.93] | Yb<br>[173.05] | Lu<br>[174.97] |              |              |              |              |             |             |    |    |    |    |    |    |    |    |  |
| 9                              | Th<br>[232]    | Pa<br>[231]    | Np<br>[237]    | Pu<br>[244]    | Am<br>[243]    | Cm<br>[247]    | Bk<br>[247]    | Cf<br>[251]    | Es<br>[252]    | Fm<br>[257]    | Mn<br>[288]    | Ni<br>[289]    | Pu<br>[289]    | At<br>[210]    | Fr<br>[223]  |              |              |              |             |             |    |    |    |    |    |    |    |    |  |

The only metal that is liquid at room temperature.

Period 3  
element with  
7 valence  
electrons.

Semi-conductor  
metalloid used  
in electronics.

Only one valence electron, but not part of a family.

Extremely reactive non-metal.

Group that will form 2+ cations.

The only non-metal that is liquid at room temperature.

Completely  
un-reactive family  
with complete  
valence shell.

Extremely reactive metal.

Group with 4 valence electrons.

# Periodic Trends Review Answers

**LL Leslie**



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