Graphing Periodic Trends Lab

Name_____Hr___

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

ionization energy - The energy required to remove an electron from an atom.

DATA

Properties of the First 18 Elements

Properties of the First 18		Elements	
Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy
н	1	0.79	314
He	2	0.49	567
Li	3	2.05	124
Ве	4	1.4	215
8	5	1.17	191
C	6	0.91	260
M	7	0.75	335
0	8	0.65	314
F	9	0.57	402
No	10	7	2
Ma	1.1	2.23	119
Mg	12	1.72	176
A.I	13	1.82	138
Si	1.4	1.46	188
P	15	1.23	242
S	16	1.09	239
CI	17	0.97	299
Ar	18	?	2

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with guestions marks.

- Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)
- 2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)

Periodic Trends Lab Answers

Daniel F McAuley

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