

Unit 2: Atomic Scientists and Atomic Structure

Unit II. LT # 1&2

Atomic Scientists and Atomic Structure Magic Squares

Directions: Put the number of the definition from the list below into the square with the appropriate term. Check your answers by adding the numbers to see if all the sums of all rows, both across and down add up to the same number, the Magic #.

Democritus <u>2</u>	John Dalton <u>7</u>	J.J. Thomson <u>18</u>	Henry Moseley <u>12</u>	Total <u>39</u>
Ernest Rutherford <u>8</u>	Proton <u>5</u>	Atom <u>11</u>	Henri Becquerel <u>15</u>	<u>39</u>
Millikan <u>13</u>	Neutron <u>17</u>	Nucleus <u>6</u>	Alpha Particle <u>3</u>	<u>39</u>
Electron <u>16</u>	Model <u>10</u>	Mendeleev <u>4</u>	Electron Cloud <u>9</u>	<u>39</u>
Total <u>39</u>	<u>39</u>	<u>39</u>	<u>39</u>	

Magic Number: 39

1. Represented by a symbol; all are found on the Periodic Table
2. Made a mental model of the atom; Greek philosopher
3. Used by Rutherford in his experiment; made of two protons and two neutrons (aka: He)
4. Considered to be the father of the periodic table.
5. The positive particle in the nucleus of an atom. (Not a fundamental particle)
6. The tiny positive core of an atom; contains protons and neutrons.
7. Credited with the first atomic theory model of the atom which had four main postulates.
8. Discovered the nucleus using his gold foil experiment.
9. Current explanation of where electrons might be found in the atom.
10. Used by scientists to explain something we can not see or understand.
11. The smallest particle of an element that has the properties of that element.
12. Arranged the periodic table in terms of the atomic number of the element.
13. Determined the exact charge of the electron with his oil drop experiment.
14. Mass of protons and neutrons
15. Along with Marie and Pierre Curie, discovered radioactivity.
16. The negative particle that circles the nucleus

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J Spring



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