

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:-

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- Based on Moseley law: $\sqrt{\lambda} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e^- configuration.

Transition elements; they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G_1 and $G_2 = ns^1-2$

p-Block = $G_{13} - G_{18} = ns^2 np^1-6$

d-Block = $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid = $2 = 5f - 7$

actinoids = $2 = 5f - 10$
 G_3 and P_6

Diagonal Relation;

Li → Be → B → similar
Na → Mg → Al → change ratio
size

Nomenclature;

0 → nil	5 → pent	Suffix = turn
1 → un	6 → hex	100 unnilpentium
2 → bi	7 → sept	
3 → tri	8 → oct	
4 → quad	9 → enn	

EFFECTIVE NUCLEAR CHARGE

$$Z_{\text{eff}} = Z - \sigma$$

Trends - Period → $Z_{\text{eff}} \uparrow$

Group → Z_{eff} almost remains same. or $Z_{\text{eff}} \downarrow$

Shielding → $s > p > d > f$

→ more +ve charge, more is Z_{eff}

→ more -ve charge, less is Z_{eff}

IONISATION ENERGY

- Energy required to remove an electron from an isolated gaseous atom.



$$IE = IE_1 < IE_2 < IE_3 \dots$$

Trends - $Z_{\text{eff}} \propto IE$

Period - $Z_{\text{eff}} \uparrow = IE \uparrow$

Group - $Z_{\text{eff}} \downarrow = IE \downarrow$

$G_2 = O > AL > Ga > In$

$G_{14} = C > Si > Ge > Pb > Sn$

Period 2:-



ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

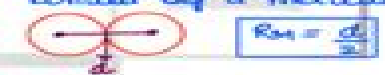
Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covalent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends; Atomic Radius $\propto \frac{1}{Z_{\text{eff}}}$

Period → $Z_{\text{eff}} \uparrow = \text{Atomic Radius} \downarrow$
Group → $Z_{\text{eff}} \downarrow = \text{Atomic Radius} \uparrow$

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

- Energy required or released when an e^- is added to outermost orbit of an isolated gaseous atom. It is known as Electron gain enthalpy.



Note - In majority cases $\Delta_{\text{egH}} 1$ is -ve (exothermic)

Exception = $\begin{matrix} B \\ N \end{matrix}$ } $\Delta_{\text{egH}} 1 = +ve$
Noble gas } due to stable configuration

Note - Successive electron gain enthalpy is always +ve.

Trends; $\Delta_{\text{egH}} \propto Z_{\text{eff}}$

Period - $Z_{\text{eff}} \uparrow \Rightarrow \Delta_{\text{egH}} \uparrow$

Group - $Z_{\text{eff}} \downarrow \Rightarrow \Delta_{\text{egH}} \downarrow$

$G_2 = S > Se > Te > Po > O$
→ due to small size Reason!
→ e^-e^- repulsion.

Periodic Classification 3rd Chapter 11th Class

Disha Experts

A decorative red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the 'Disha Experts' text.

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The novel begins with Richard Papan, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

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