

Periodic Trends Worksheet

Using a few minutes before today's activity, complete the following questions. Please post your answers on the board when you are done.

Additional help on the periodic trends can be found on this video: <https://www.youtube.com/watch?v=0D8Wp-2GgY0>

1. What does electronegativity measure and how does it change across the periodic table? (electronegativity is largest) in which direction, left-to-right, up-to-down, and bottom-to-top?
2. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
3. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. Fluorine > Li
 - b. Li > Fluorine
 - c. Cl > Cs
4. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. R > Cs
 - b. Cs > R
 - c. Li > R
5. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. Cl > Cs
 - b. Fluorine > Li
 - c. R > R
6. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. Fluorine > Cs
 - b. R > R
 - c. R > R
7. Ionization energy measures how a larger atomic radius affects the atom's ability to lose an electron.
 - a. R > R
 - b. Cs > Cs
 - c. Fluorine > R
8. Which is the most electronegative element on the periodic table? (Fluorine)

Periodic Trends Questions And Answers

Sasha Hu



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