

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 11 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

*Groups 1 to 10 are named after their first members i.e. Group 1 is the scandium group

1	2											13	14	15	16	17	18
1 H 1.01 Hydrogen	2 He 4.003 Helium											3 B 10.81 Boron	4 C 12.01 Carbon	5 N 14.01 Nitrogen	6 O 16.00 Oxygen	7 F 19.00 Fluorine	8 Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.012 Beryllium											9 Al 26.98 Aluminum	10 Si 28.09 Silicon	11 P 30.97 Phosphorus	12 S 32.06 Sulfur	13 Cl 35.45 Chlorine	14 Ar 39.95 Argon
5 Na 22.99 Sodium	6 Mg 24.31 Magnesium											15 Ga 69.72 Gallium	16 Ge 72.64 Germanium	17 As 74.92 Arsenic	18 Se 78.96 Selenium	19 Br 79.90 Bromine	20 Kr 83.80 Krypton
7 K 39.10 Potassium	8 Ca 40.08 Calcium											21 In 114.82 Indium	22 Sn 118.71 Tin	23 Sb 121.76 Antimony	24 Te 127.60 Tellurium	25 I 126.91 Iodine	26 Xe 131.29 Xenon
9 Rb 85.47 Rubidium	10 Sr 87.62 Strontium											27 Tl 204.38 Thallium	28 Pb 207.2 Lead	29 Bi 208.98 Bismuth	30 Po 209 Polonium	31 At 210 Astatine	32 Rn 222 Radon
11 Cs 132.91 Cesium	12 Ba 137.33 Barium											33 Hg 200.59 Mercury	34 Tl 204.38 Thallium	35 Pb 207.2 Lead	36 Bi 208.98 Bismuth	37 Po 209 Polonium	38 At 210 Astatine
13 Fr 223 Francium	14 Ra 226 Radium											39 Au 196.97 Gold	40 Pt 195.08 Platinum	41 Ag 107.87 Silver	42 Cd 112.41 Cadmium	43 In 114.82 Indium	44 Sn 118.71 Tin
												45 Rh 101.07 Rhodium	46 Pd 106.37 Palladium	47 Cu 63.55 Copper	48 Zn 65.38 Zinc	49 Ga 69.72 Gallium	50 Ge 72.64 Germanium
												51 Ir 223.03 Iridium	52 Pt 195.08 Platinum	53 Au 196.97 Gold	54 Hg 200.59 Mercury	55 Tl 204.38 Thallium	56 Pb 207.2 Lead
												57 La 138.91 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.91 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.91 Promethium	62 Sm 150.36 Samarium
												63 Eu 151.96 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.93 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.93 Holmium	68 Er 167.26 Erbium
												69 Tm 168.93 Thulium	70 Yb 173.05 Ytterbium	71 Lu 174.96 Lutetium	72 Hf 178.49 Hafnium	73 Ta 180.95 Tantalum	74 W 183.84 Tungsten
												75 Re 186.21 Rhenium	76 Os 190.23 Osmium	77 Ir 192.22 Iridium	78 Pt 195.08 Platinum	79 Au 196.97 Gold	80 Hg 200.59 Mercury
												81 Tl 204.38 Thallium	82 Pb 207.2 Lead	83 Bi 208.98 Bismuth	84 Po 209 Polonium	85 At 210 Astatine	86 Rn 222 Radon
												87 Fr 223 Francium	88 Ra 226 Radium	89 Ac 227 Actinium	90 Th 232.04 Thorium	91 Pa 231.04 Protactinium	92 U 238.03 Uranium
												93 Np 237.05 Neptunium	94 Pu 239.05 Plutonium	95 Am 243.06 Americium	96 Cm 247.07 Curium	97 Bk 247.07 Berkelium	98 Cf 251.08 Californium
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												105 Db 262.11 Dubnium	106 Sg 266.12 Seaborgium	107 Bh 264.10 Bohrium	108 Hs 277.13 Hassium	109 Mt 268.10 Meitnerium	110 Ds 271.10 Darmstadtium
												111 Rg 272.11 Roentgenium	112 Cn 285.10 Copernicium	113 Nh 284.10 Nihonium	114 Fl 289.10 Flerovium	115 Mc 288.10 Moscovium	116 Lv 293.10 Livermorium
												117 Ts 294.10 Tennessine	118 Og 294.10 Oganesson	119 Nh 294.10 Nihonium	120 Fl 294.10 Flerovium	121 Mc 294.10 Moscovium	122 Lv 294.10 Livermorium

CLASSIFICATION

Metals	Nonmetals
<div>Active</div> <div>Pre-transition</div> <div>Alkali earth</div> <div>Alkaline</div>	<div>Chalcogens</div>
<div>Transition</div>	<div>Intermediate</div>
<div>Post-transition</div>	<div>Metalloids</div>
<div>None (Ru-Pt, Os-Au)</div>	<div>None gas</div>

Active metals

Pre-transition

Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths

Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides

Soft, dense and reactive metals, those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals

Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalyst properties. Very weakly electropositive; inert to acids.

Frontier metals

Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and tellurates.

Metals

Nonmetals

Metalloids

Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionization energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals

H and N are colorless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionization energies; low to high electron affinities; moderate to high electronegativity values; relatively poor to moderately strong oxidizing agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (P₂O, CO, NO, N₂O).

*as black P, the most stable form in ambient conditions

Group 1

This starts as Sc-P, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lr. Trends going down Sc-P, and then La branch, resemble those of groups 1–2.

• Lu branch, more resemble those of groups 1–2

• Lu branch, more resemble those of groups 1–2

Radioactive

Approximation is likely to be either a metastable or a good oxidation agent

Periodic Table Guide Answers

Daniel F McAuley

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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