

Percent Yield Practice

1. Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.



$$\frac{24.8 \text{ g CaCO}_3}{1} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaCO}_3} \cdot \frac{56.08 \text{ g CaO}}{1 \text{ mol CaO}} = 13.9 \text{ g CaO} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{13.1}{13.9} \cdot 100 = \boxed{94.24\%} \Rightarrow \boxed{94.2\%}$$

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?



$$\frac{1.87 \text{ g Al}}{1} \cdot \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \cdot \frac{159.62 \text{ g CuSO}_4}{1 \text{ mol CuSO}_4} = 16.6 \text{ g CuSO}_4 \quad \text{Used}$$

$$\frac{9.65 \text{ g CuSO}_4}{159.62 \text{ g CuSO}_4} \cdot \frac{1 \text{ mol CuSO}_4}{1 \text{ mol CuSO}_4} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol CuSO}_4} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 3.84 \text{ g Cu}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{3.65}{3.84} \cdot 100 = \boxed{95.05\%} \quad \text{TY} \rightarrow \boxed{95.1\%}$$

3. If you start with 5.0 g of C_3H_8 and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?



$$\frac{5.0 \text{ g C}_3\text{H}_8}{44.11 \text{ g C}_3\text{H}_8} \cdot \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 8.2 \text{ g H}_2\text{O} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100$$

$$\frac{\text{PY} \cdot \text{TY}}{100} = \text{AY} = \frac{75 \cdot 8.2}{100} = \boxed{6.2 \text{ g H}_2\text{O}} \quad \text{AY}$$

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?



$$\frac{20 \text{ g CaO}}{56.08 \text{ g CaO}} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaO}} \cdot \frac{1 \text{ mol Na}_2\text{O}}{1 \text{ mol CaO}} \cdot \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{20 \text{ g Na}_2\text{O}}$$

Percent Yield Answer Key

LP Steffe

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LP Steffe' text.

Percent Yield Answer Key:

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ACCA P4 - Advanced Financial Management - Study Text 2013 BPP Learning Media,2011-12-15 The BPP Study Text provides a comprehensive treatment of the updated ACCA syllabus for P4 It addresses all learning outcomes and the higher skills required in an integrated and practical way The material despite the technical nature of certain areas follows a practical common sense approach with plenty of case studies and real life business examples The key points of each topic are summarised in a chapter roundup and tested in a diagnostic quiz at the end of each chapter A question bank at the end of the book provides practice on exam style questions

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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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Percent Yield Answer Key Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Percent Yield Answer Key**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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