

Orgo 2OG3-Cheat Sheet

	Strong base Weak nucleophile	Strong base Strong nucleophile	Weak base Strong nucleophile	Weak base Weak nucleophile
1°	E2	E2 SN2	SN2	
2°	E2	E2 SN2	SN2	
3°	E2	E2	SN1	SN1 E1

Strong base Weak nucleophile	Strong base Strong nucleophile	Weak base Strong nucleophile	Weak base Weak nucleophile
NaH DBN DBU	HO ⁻ MeO ⁻ EtO ⁻	I ⁻ Br ⁻ Cl ⁻ RS ⁻ HS ⁻ RSH H ₂ S	H ₂ O MeOH EtOH

-SN1 & E1 only seen for 1° & 2° C's when carbocation is resonance stabilized

Predicting Products:

1. Substrate
-1°C, 2°C, 3°C
2. Reagent
-Nucleophilicity/Basicity
3. Stereochemistry
SN2 → Inversion
SN1 → Racemic
E1/E2 → Trans > Cis
4. Regioselectivity
SN1/SN2 → α-carbon
E2 → Zaitsev favoured
-Unless base is hindered
E1 → Zaitsev favoured always

E2: Base deprotonates, halide leaves & π-bond forms
-More exposed adjacent C protons the better
3°C > 2°C > 1°C

Regioselectivity

- Greater substituted alkene increases stability
- Increasing bulky nucleophiles prefer less substituted

Stereoselectivity

- Trans alkenes > stability than cis alkenes

Stereospecificity

- Adjacent proton must be anti to leaving group
- Opposite stereochemistry (Dash-Wedge)

SN2: Nucleophile attacks electrophile → Dissociates leaving group
Nucleophile strength → Charge > Basicity > Polarizability
-Back side attack therefore inverse configuration about C atom (S → R)
-Steric bulk prohibits attack
1°C > 2°C > 3°C

- Bulky adjacent carbons disallow interaction (Ex. Neopentyl)
- Bulky nucleophiles also not favoured for SN2

Effect of solvents → Refer loose-leaf

E1: After ionization of alkyl halide, deprotonation occurs & π-bond forms
-Carbocation intermediate (Unimolecular)

Regioselectivity

- Greater substituted alkene increases stability
- Trans alkenes > stability than cis alkenes

SN1: Alkyl halides undergo ionization in a polar solvent
-LG leaves, nucleophile attacks carbocation
-Carbocation intermediate (Unimolecular)

Carbocation rearrangement

- Methyl/hydride shifts to stabilize the charge

Stereochemistry

- Attack can occur from two equal paths
- Forms a racemic mixture (Opposite configurations)

Hydrohalogenation:

- Addition of H-X across a nucleophilic π-bond

Markovnikov's Rule:

- H adds to carbon with most H's (Carbocation stability)

Anti-Markovnikov Addition:

- In Presence of peroxides → Radical's form and H adds to carbon with less H's

Stereochemistry

- If chiral center formed → Both enantiomers formed (Racemic)

Acid-Catalyzed Hydration (H₂O Addition):

- Addition of H₂O across a double bond in presence of dilute acid



Hydrohalogenation & Hydration Both Involve a Carbocation:

- Have chance of undergoing rearrangement



- If we don't want the rearrangement products

-Oxymercuration & Demercuration

- Follows Markovnikov's rule

Rate of reaction:

- 3°C > 2°C > 1°C (Carbocation stability)

Alcohol Dehydration:

- Alcohols can be dehydrated using conc. H₂SO₄



Orgo Mechanisms Cheat Sheet

John T. Moore, Richard H. Langley



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Table of Contents Orgo Mechanisms Cheat Sheet

1. Understanding the eBook Orgo Mechanisms Cheat Sheet
 - The Rise of Digital Reading Orgo Mechanisms Cheat Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Orgo Mechanisms Cheat Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Orgo Mechanisms Cheat Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Orgo Mechanisms Cheat Sheet
 - Personalized Recommendations
 - Orgo Mechanisms Cheat Sheet User Reviews and Ratings
 - Orgo Mechanisms Cheat Sheet and Bestseller Lists
5. Accessing Orgo Mechanisms Cheat Sheet Free and Paid eBooks
 - Orgo Mechanisms Cheat Sheet Public Domain eBooks
 - Orgo Mechanisms Cheat Sheet eBook Subscription Services
 - Orgo Mechanisms Cheat Sheet Budget-Friendly Options
6. Navigating Orgo Mechanisms Cheat Sheet eBook Formats

- ePub, PDF, MOBI, and More
- Orgo Mechanisms Cheat Sheet Compatibility with Devices
- Orgo Mechanisms Cheat Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Orgo Mechanisms Cheat Sheet
 - Highlighting and Note-Taking Orgo Mechanisms Cheat Sheet
 - Interactive Elements Orgo Mechanisms Cheat Sheet
- 8. Staying Engaged with Orgo Mechanisms Cheat Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Orgo Mechanisms Cheat Sheet
- 9. Balancing eBooks and Physical Books Orgo Mechanisms Cheat Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Orgo Mechanisms Cheat Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Orgo Mechanisms Cheat Sheet
 - Setting Reading Goals Orgo Mechanisms Cheat Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Orgo Mechanisms Cheat Sheet
 - Fact-Checking eBook Content of Orgo Mechanisms Cheat Sheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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