

ORGANIC CHEMISTRY

FUNDAMENTALS

Types of Organic Compounds

Organic chemistry is the study of natural and synthetic materials that have carbon atoms as the key chemical feature. There are more than one million known organic compounds.

HYDROCARBON

ALKANE
 $\text{C}-\text{H}$
 $\text{C}-\text{C}$
 • Ethane: C_2H_6
 • Methyl (Me): $-\text{CH}_3$
 • Ethyl (E): $-\text{C}_2\text{H}_5$

ALKENE

• Ethene: C_2H_4
 • Diene: Two $\text{C}=\text{C}$
 • Triene: Three $\text{C}=\text{C}$

ALKYNE

• Ethyne: C_2H_2

AROMATIC

• Benzene: C_6H_6
 • Aro: C_6H_5 (Ar)

-O- ADDED

ALCOHOL

• Methanol: $\text{Me}-\text{OH}$ (methyl alcohol)
 • Phenol: $\text{Ar}-\text{OH}$
 • Diol/Glycol: (2 $-\text{OH}$)
 • Glycol: (3 $-\text{OH}$)

ETHER

• Ethoxyethane: $\text{Et}-\text{O}-\text{Et}$ (diethyl ether)

EPOXY

• Oxirane: $\text{C}_2\text{H}_4\text{O}$ (ethylene oxide)

PEROXIDE

• Acetone Peroxide: $\text{C}_4\text{H}_8\text{O}_5$

• Benzoyl Peroxide: $\text{C}_{14}\text{H}_{18}\text{O}_5$

>C=O ADDED

ALDEHYDE

• Methanal: H_2CO (formaldehyde)
 • Benzaldehyde: $\text{Ar}-\text{CHO}$

KETONE

• 2-propanone: $\text{Me}-\text{CO}-\text{Me}$ (dimethyl ketone, acetone)
 • Diketone: $\text{R}-\text{CO}-\text{R}'-\text{CO}-\text{R}''$

>COO ADDED

CARBOXYLIC ACID

• Ethanoic Acid: $\text{Me}-\text{COOH}$ (acetic acid)
 • Acetate Ion: $\text{Me}-\text{COO}^-$
 • Benzoic Acid: $\text{Ar}-\text{COOH}$

Dicarboxylic Acid

$\text{HOOC}-\text{R}-\text{COOH}$

ESTER

• Ethyl Acetate: $\text{Me}-\text{CO}-\text{OEt}$
 Other derivatives:
 • Peroxyacid: $\text{R}-\text{CO}-\text{OOH}$
 • Acid Anhydrides: $\text{RCO}-\text{O}-\text{CO}-\text{R}'$

NITROGEN ADDED

AMINE

• Methyl Amine: $\text{H}_3\text{C}-\text{NH}_2$
 • Phenyl amine: $\text{Ar}-\text{NH}_2$ (aniline)
 • $\text{R}-\text{NH}_2$ (1°), $\text{RR}'\text{NH}$ (2°), $\text{RR}'\text{R}''\text{N}$ (3°)

NITRO

$\text{R}-\text{NO}_2$

DIAZO

$\text{R}-\text{N}=\text{N}$

NITRILE

$\text{R}-\text{C}\equiv\text{N}$

• Methane Nitrile: $\text{Me}-\text{CN}$

AMIDE

• Acetamide: $\text{Me}-\text{CO}-\text{NH}_2$

SULFUR ADDED

• Thiol: $\text{R}-\text{SH}$
 • Thioether: $\text{R}-\text{S}-\text{R}'$
 • Disulfide: $\text{R}-\text{S}-\text{S}-\text{R}'$
 • Thio Ester: $\text{R}-\text{CO}-\text{SR}'$
 • Sulfide: $\text{R}-\text{SO}-\text{R}'$
 • Sulfone: $\text{R}-\text{SO}_2-\text{R}'$
 • Sulfonic Acids: $\text{R}-\text{SO}_3\text{H}$

HALOGEN ADDED

• Haloalkane: $\text{Me}-\text{Cl}$ chloromethane
 • Halohydrate: $\text{Ar}-\text{X}$
 • Chloroalkane: $\text{Ar}-\text{Cl}$
 • Acyl Halide: $\text{R}-\text{CO}-\text{X}$
 • Aryl Halide: $\text{Ar}-\text{X}$

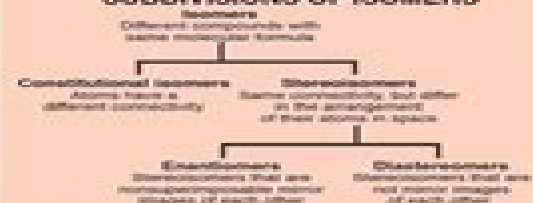
Formulas & Isomers

- Molecular Formula:** Elemental symbols with subscripts denote the composition of a compound
- Empirical Formula:** Subscripts denote the relative elemental composition
- Graphical depiction**
 - Dash Formula:** Diagram all atoms, show bonds as dashes
 - Bond-Line Formula:** Hide H; carbon atoms are depicted as lines, but other atoms are shown explicitly



- Newman Projection:** 2-d depiction
- 3-dimensionals:** Wedges of sawhorse denote structure

SUBDIVISIONS OF ISOMERS



- Constitutional Isomers:** Different bonding connectivity (e.g., rings, bonds, branching, substituent positions)
- Tautomers:** Easily interconverted structural isomers (e.g., keto-enol for ketones)
- Chiral:** Not identical with its mirror image
- Achiral:** Has a plane of symmetry (i.e., superimposable on its mirror image)
- Epimers:** Pair of diastereomers which differ only in the configuration of one atom
- More than 1 chiral center
 - n chiral centers, $\leq 2^n$ stereoisomers
 - Meso: 2 chiral centers, 4 isomers: 3 stereoisomers, 1 achiral (mirror-plane)

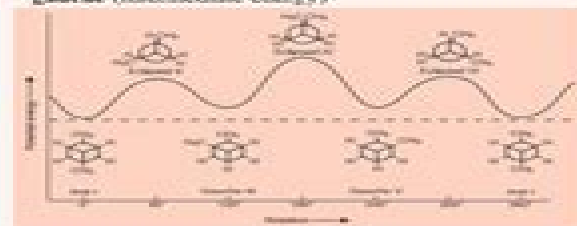


- R/S Notation:** The 4 different atoms or groups attached to a central atom are ranked (a), (b), (c), and (d), by molar mass. The lowest (d) is directed away from the viewer and the sequence of a-b-c produces clockwise (R) or counter-clockwise (S) configuration
- Chiral (Optically Active):** + or - the rotation of plane polarized light, R/S: opposite effects

- Racemic:** 50/50 mixture of stereoisomers (no net optical activity)
- Nomenclature:** Note R/S and +/- in the compound name (e.g., R (+) bromochloromethanol)
- Fisher-Projection:** Diagram depicts chiral/3-D structure
- Molecular Conformations:** Molecule exhibits structural variation due to free rotation about C-C single bond



- Newman Diagram:** Depict rotation about a C-C bond; eclipsed (high energy), anti (low energy), and gauche (intermediate energy)



Common Terms

Aliphatic: Non-aromatic
Aromatic: Benzene ring
Conjugation: Sequence of alternating double (or triple) and single bonds
Dielectric Effect: Polar solvent stabilizes ion formation
Exergonic Reaction: Chemical reaction

that loses energy during the process of the reaction
Heterocyclic: Non-carbon atom in the ring structure
Hydrocarbon: Compound of H and C
Ketone: Molecule made up of a monosaccharide bonded to a ketone

Monosaccharide: Carbohydrate that cannot be reduced by hydrolysis into another simple sugar
Olefin: Alkene
Paraffin: Alkane
Radioisotope: Isotope that is radioactive

Saturated: Maximum # of Hs (all C-C single bonds)
Tetrose: Monosaccharide with four carbon atoms
Unsaturated: At least one C-C multiple bond
Zymology: The study of the process of fermentation

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Table of Contents Organic Chemistry Fundamentals Study Guide

1. Understanding the eBook Organic Chemistry Fundamentals Study Guide
 - The Rise of Digital Reading Organic Chemistry Fundamentals Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Fundamentals Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Fundamentals Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Fundamentals Study Guide
 - Personalized Recommendations
 - Organic Chemistry Fundamentals Study Guide User Reviews and Ratings
 - Organic Chemistry Fundamentals Study Guide and Bestseller Lists
5. Accessing Organic Chemistry Fundamentals Study Guide Free and Paid eBooks
 - Organic Chemistry Fundamentals Study Guide Public Domain eBooks
 - Organic Chemistry Fundamentals Study Guide eBook Subscription Services
 - Organic Chemistry Fundamentals Study Guide Budget-Friendly Options

6. Navigating Organic Chemistry Fundamentals Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Organic Chemistry Fundamentals Study Guide Compatibility with Devices
 - Organic Chemistry Fundamentals Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Fundamentals Study Guide
 - Highlighting and Note-Taking Organic Chemistry Fundamentals Study Guide
 - Interactive Elements Organic Chemistry Fundamentals Study Guide
8. Staying Engaged with Organic Chemistry Fundamentals Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Fundamentals Study Guide
9. Balancing eBooks and Physical Books Organic Chemistry Fundamentals Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Fundamentals Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry Fundamentals Study Guide
 - Setting Reading Goals Organic Chemistry Fundamentals Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chemistry Fundamentals Study Guide
 - Fact-Checking eBook Content of Organic Chemistry Fundamentals Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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