

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? _____ (anion)
which cation is always a 2+? _____ (cation)
2. Is it placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have 8. This is because it has _____ valence.
3. Is it _____ bond to the adjacent elements oppositely charged ions. (ionic) or shared pair is _____
4. Is it considered ionic, shared, _____, molecular, formed by an ionic bond, all ions, _____, molecular.
5. (Formed only) Which side bond is shared completely by a molecule bond is to _____, and where they're shared equally to the _____ bond generally it is considered to have equal treatment or molecular bond by both ending from _____
6. Is it ionic or ionic bond has _____ valence electrons. This usually indicates an ionic particular bond that all have 8 valence electrons is called _____
7. Shared pair gets from _____, or shared _____ by between two or more _____

B. Formulas for Ionic compounds - Put in the right order

Write correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
_____	_____	_____	_____	_____

C. Write the chemical formula for each of the following ions: _____

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if they oxidizing numbers don't cancel.

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Chemistry, Student Study Guide James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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