EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES (CHEMISTRY)

TABLE 1 PHYSICAL CONSTANTS

| NAME | SYMBOL | VALUE |
|---------------------------------|----------------|---|
| Magnitude of charge on electron | e | 1,6 × 10 ⁻¹⁹ C |
| Mass of an electron | m _e | 9,1 × 10 ⁻³¹ kg |
| Standard pressure | p* | 1,01 × 10 ⁵ Pa |
| Molar gas volume at STP | V _m | 22,4 dm ³ ·mol ⁻¹ |
| Standard temperature | T* | 273 K |
| Avogadro's constant | N _A | 6,02 × 10 ²³ mol ⁻¹ |
| Faraday's constant | F | 96 500 C · mol-1 |

TABLE 2 CHEMISTRY FORMULAE

$$n = \frac{m}{M}$$

$$n = \frac{N}{N_A}$$

$$n = \frac{V}{V_m}$$

$$G = \frac{n}{V} \text{ OR } G = \frac{m}{MV}$$

$$K_w = \begin{bmatrix} H_3O^+ \end{bmatrix} \cdot \begin{bmatrix} OH^- \end{bmatrix} = 1 \times 10^{-14}$$

$$\text{at 25 °C (298 K)}$$

$$Q = It$$

$$Q = nF$$

$$E_{cell}^\theta = E_{oxidising agent}^\theta - E_{reducing agent}^\theta$$

$$E_{cell}^\theta = E_{oxidising agent}^\theta - E_{reducing agent}^\theta$$

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Table of Contents Physical Science P2 Grade1exampler

- 1. Understanding the eBook Physical Science P2 Grade1exampler
 - The Rise of Digital Reading Physical Science P2 Grade1exampler
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science P2 Grade1exampler
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science P2 Grade1exampler
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science P2 Grade1exampler
 - Personalized Recommendations
 - Physical Science P2 Grade1exampler User Reviews and Ratings
 - Physical Science P2 Grade1exampler and Bestseller Lists

- 5. Accessing Physical Science P2 Grade1exampler Free and Paid eBooks
 - Physical Science P2 Grade1exampler Public Domain eBooks
 - Physical Science P2 Grade1exampler eBook Subscription Services
 - Physical Science P2 Grade1exampler Budget-Friendly Options
- 6. Navigating Physical Science P2 Grade1exampler eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Physical Science P2 Grade1exampler Compatibility with Devices
 - Physical Science P2 Grade1exampler Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science P2 Grade1exampler
 - Highlighting and Note-Taking Physical Science P2 Grade1exampler
 - Interactive Elements Physical Science P2 Grade1exampler
- 8. Staying Engaged with Physical Science P2 Grade1exampler
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science P2 Grade1exampler
- 9. Balancing eBooks and Physical Books Physical Science P2 Grade1exampler
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science P2 Grade1exampler
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science P2 Grade1exampler
 - Setting Reading Goals Physical Science P2 Grade1exampler
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science P2 Grade1exampler
 - Fact-Checking eBook Content of Physical Science P2 Grade1exampler
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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