

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

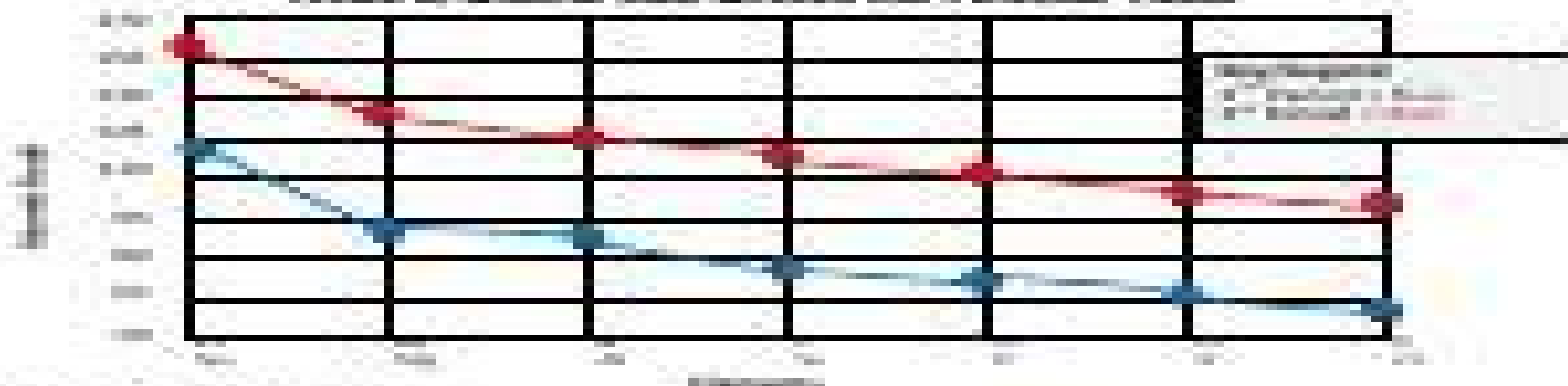
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property. Sketch it's data points (graph paper)
2. Is a slight downward trend in this graph Period? If not, fill in the other graphs
3. Label the box as best matches ionization
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9F	10Ne	11Na	12Mg	13Al	14Si	15P	16S
3	17Cl	18Ar	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Ba	56La

Now label the properties above in the table below:

Trend in Atomic Size Across the Periodic Table



Observed Trend Questions

1. Following the trend in atomic size across the periodic table: decreases
2. This means two atoms that are more metallic possess a difference that gives all elements, making themselves, almost an insulating material.

Expected Trend Questions

1. Theoretical trend of electronegativity: a measure of the ability of an atom to attract electrons. It is a measure of the periodic table's metallic
2. This concept between them are more metallic and non-metallic properties/differences (the more all elements) making them, otherwise, insulating.

Periodicity Activity Okstate Answer Key

Robert R. Redfield



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The Journal of the Oklahoma State Medical Association ,1991

Ecology Abstracts ,1988

Agrindex ,1993

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