

Answer KEY

Unit 3 Outline/ Study Guide - Electron Configuration and Periodic Trends

Essential Skills/ State Standards:

1. Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
2. Students know how to use the periodic table to identify metals, semimetals, nonmetals, halogens, alkali metals, alkaline earth metals, and transition metals.
3. Students know how to use the periodic table to identify the following trends:
 ionization energy, electronegativity, and the relative sizes of ions and atoms
 (this includes knowing the definitions of ionization energy, electronegativity, cation, and anion.)

Study Guide - (use your notes and worksheets for a more complete review)

1. Identify the following things on the periodic table:

-metals	-alkali metals
-metalloids (semi-metals)	-alkaline earth metals
-transition metals	-halogens
-nonmetals	-noble gases
-Periods	-Groups/ families
Identify groups with the following numbers of valence electrons: 1, 2, 3, 4, 5, 6, 7, 8	

2. Why do the elements in the groups (columns) behave so similar to each other even if they have great differences in their number of protons? **same number of valence electrons and similar chemical properties.**
3. What is the connection between the following terms: **octet rule, valence electrons, and ions?**
In order to achieve the octet rule, atoms gain or lose their valence electrons. They become charged ions when this occurs.
4. a. List the elements that are exceptions to the octet rule b) Why don't these atoms achieve the octet rule?
 a. Small atoms are happy with 2 electrons (ex: H, He, Li, Be, B)
 b. because the 1s energy level only holds 2 electrons

Periodic Trends Test Review Key

Scott C. Dulebohn

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Table of Contents Periodic Trends Test Review Key

1. Understanding the eBook Periodic Trends Test Review Key
 - The Rise of Digital Reading Periodic Trends Test Review Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Test Review Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Test Review Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Test Review Key
 - Personalized Recommendations
 - Periodic Trends Test Review Key User Reviews and Ratings
 - Periodic Trends Test Review Key and Bestseller Lists
5. Accessing Periodic Trends Test Review Key Free and Paid eBooks

- Periodic Trends Test Review Key Public Domain eBooks
- Periodic Trends Test Review Key eBook Subscription Services
- Periodic Trends Test Review Key Budget-Friendly Options
- 6. Navigating Periodic Trends Test Review Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Test Review Key Compatibility with Devices
 - Periodic Trends Test Review Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Test Review Key
 - Highlighting and Note-Taking Periodic Trends Test Review Key
 - Interactive Elements Periodic Trends Test Review Key
- 8. Staying Engaged with Periodic Trends Test Review Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Test Review Key
- 9. Balancing eBooks and Physical Books Periodic Trends Test Review Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Test Review Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Test Review Key
 - Setting Reading Goals Periodic Trends Test Review Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Test Review Key
 - Fact-Checking eBook Content of Periodic Trends Test Review Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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