

CHAPTER 5: PERIODIC

The Periodic Law

SECOND PART

Student's Assistant Answered the following questions for the second part of the lab.

1. If you consider the groups above with 13 protons, and the elements below with 14 protons, answer the following questions:

- _____ a. The first is a halogen (halogens)?
- _____ b. The first is a noble gas (noble)?
- _____ c. Is the element in period 4 or is another period number observed? If so, which one? (period 4 or 5?)

2. Consider the elements with the following electron configurations for $1s^2 2s^2 2p^2$:

- _____ a. The first element above the noble gas (noble)?
- _____ b. It is a noble gas (noble)?
- _____ c. The group number will equal $10 + 2 = 12$. (Group 12?)

- _____ d. The noble gases are chemically inert, e.g., $1s^2 2s^2 2p^6$
- _____ e. The noble gases are other elements, chemical inert, e.g., $1s^2 2s^2 2p^6$

- _____ f. Explain the first element above the noble gas (noble)?

- _____ g. The first is a noble gas (noble)?
- _____ h. Is the electron configuration the same as the noble gas (noble)?

3. Write the noble gas electron configuration for the first element of Period 4 (K) and Period 5 (Rb) and the following questions are particular elements:

- _____ a. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ b. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ c. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ d. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ e. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ f. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ g. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ h. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ i. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ j. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ k. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ l. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ m. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ n. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ o. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ p. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ q. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ r. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ s. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ t. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ u. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ v. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ w. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ x. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ y. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$
- _____ z. The first element above the noble gas (noble), e.g., $1s^2 2s^2 2p^6$

Periodic Law Mixed Review Answer Key

Raffaela Di Napoli

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