

# The Periodic Law

## SECTION 2

Select the correct answer to the following questions for the space provided.

1. \_\_\_\_\_ for the modern periodic table, elements are ordered
  - (a) according to increasing atomic mass
  - (b) according to Mendeleev's original design
  - (c) according to increasing atomic number
  - (d) based on how they were discovered
2. \_\_\_\_\_ Mendeleev's original short periodic table gave useful chemical properties of elements appeared in regular intervals and also predicted elements were missing because of his saying
  - (a) theory
  - (b) prediction
  - (c) atomic number
  - (d) discovery
3. \_\_\_\_\_ The modern periodic law states that
  - (a) all new elements with the same properties are listed in the same group
  - (b) the physical and chemical properties of elements are functions of atomic number
  - (c) elements exhibit properties of both periods and groups
  - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. \_\_\_\_\_ The discovery of the noble gases changed Mendeleev's periodic table
  - (a) adding a row
  - (b) adding a new period
  - (c) adding a new column
  - (d) adding a new group
5. \_\_\_\_\_ The main reason the properties of the noble gases is that they are
  - (a) inactive
  - (b) unreactive
  - (c) large gaseous molecules
  - (d) noble
6. \_\_\_\_\_ Likewise, the first column in Group 1 has an atomic number of 3. The second element in this group has an atomic number of
  - (a) 4
  - (b) 10
  - (c) 11
  - (d) 19
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
  - (a) 10
  - (b) 18
  - (c) 20
  - (d) 22
8. How many elements are in this group?
  - (a) 10
  - (b) 18
  - (c) 20
  - (d) 22
9. What is the number of protons in this element?
  - (a) 10
  - (b) 18
  - (c) 20
  - (d) 22

# Periodic Law Chapter 5 Test Answers

**Ying Liu**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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