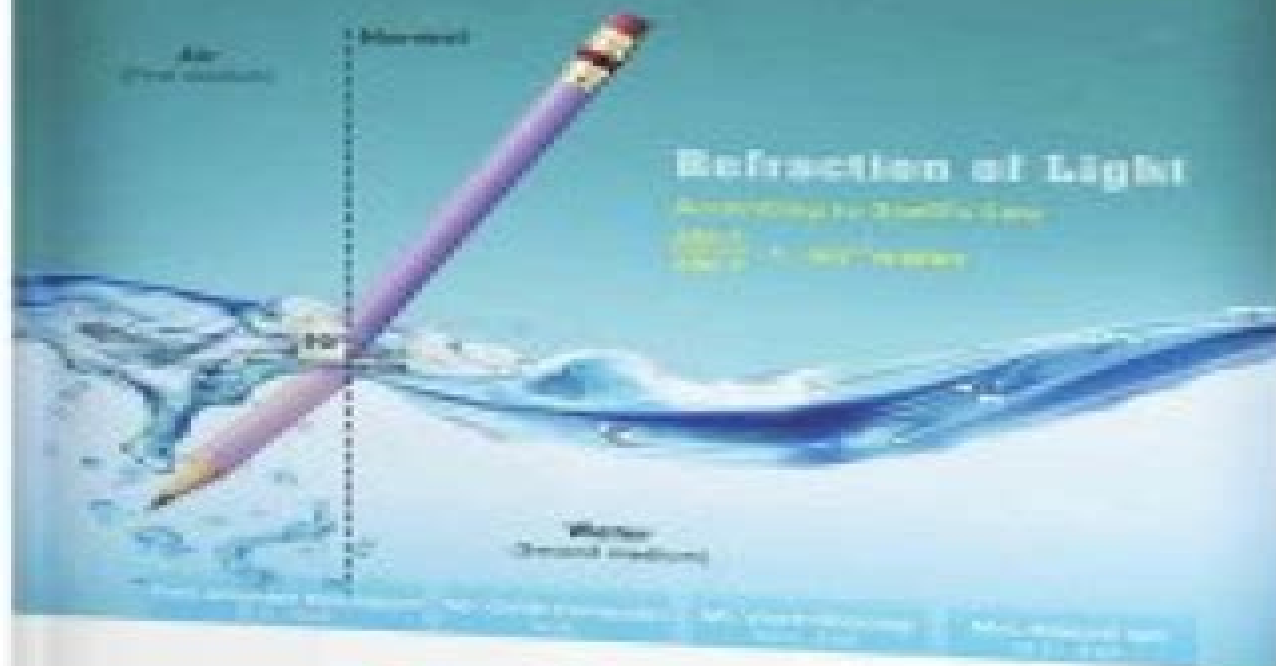


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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g/mol

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 $80 \text{ g I}_2\text{O}_5 \times \frac{1 \text{ mol I}_2\text{O}_5}{1 \text{ mol I}_2} \times \frac{1 \text{ mol I}_2}{1 \text{ mol I}_2\text{O}_5} = 33.3 \text{ g I}_2$
 $28 \text{ g CO} \times \frac{1 \text{ mol CO}}{1 \text{ mol CO}} = 1 \text{ mol CO}$

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 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$