



## Percent Composition Of Compounds

DATE \_\_\_\_\_

Find the percent composition in the following compounds.

1.  $\text{H}_2\text{O}$

$$\text{molar mass: } 2 \times 1.01 + 16.00 = 18.02 \text{ g mol}^{-1}$$

$$\% \text{H} = \frac{2 \times 1.01}{18.02} \times 100\% = 11.2\%$$

$$\% \text{O} = \frac{16.00}{18.02} \times 100\% = 88.8\%$$

2.  $\text{CO}_2$

$$\text{molar mass: } 12.01 + 2 \times 16.00 = 44.01 \text{ g mol}^{-1}$$

$$\% \text{C} = \frac{12.01}{44.01} \times 100\% = 27.3\%$$

$$\% \text{O} = \frac{2 \times 16.00}{44.01} \times 100\% = 72.7\%$$

3.  $\text{NaOH}$

$$\text{molar mass: } 22.99 + 16.00 + 1.01 = 40.00 \text{ g mol}^{-1}$$

$$\% \text{Na} = \frac{22.99}{40.00} \times 100\% = 57.5\%$$

$$\% \text{H} = \frac{1.01}{40.00} \times 100\% = 2.5\%$$

$$\% \text{O} = \frac{16.00}{40.00} \times 100\% = 40.0\%$$

4.  $\text{H}_2\text{SO}_4$

$$\text{molar mass: } 2 \times 1.01 + 32.07 + 4 \times 16.00 = 98.09 \text{ g mol}^{-1}$$

$$\% \text{H} = \frac{2 \times 1.01}{98.09} \times 100\% = 2.1\%$$

$$\% \text{S} = \frac{32.07}{98.09} \times 100\% = 32.7\%$$

$$\% \text{O} = \frac{4 \times 16.00}{98.09} \times 100\% = 65.2\%$$

5.  $\text{Ca(OH)}_2$

$$\text{molar mass: } 40.08 + 2 \times 16.00 + 2 \times 1.01 = 74.10 \text{ g mol}^{-1}$$

$$\% \text{Ca} = \frac{40.08}{74.10} \times 100\% = 54.1\%$$

$$\% \text{H} = \frac{2 \times 1.01}{74.10} \times 100\% = 2.7\%$$

$$\% \text{O} = \frac{2 \times 16.00}{74.10} \times 100\% = 43.2\%$$

# Percentage Composition Chemistry Answerkey

**Karin Nielsen-Saines**



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## **Table of Contents Percentage Composition Chemistry Answerkey**

1. Understanding the eBook Percentage Composition Chemistry Answerkey
  - The Rise of Digital Reading Percentage Composition Chemistry Answerkey
  - Advantages of eBooks Over Traditional Books
2. Identifying Percentage Composition Chemistry Answerkey
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Percentage Composition Chemistry Answerkey
  - User-Friendly Interface
4. Exploring eBook Recommendations from Percentage Composition Chemistry Answerkey
  - Personalized Recommendations
  - Percentage Composition Chemistry Answerkey User Reviews and Ratings
  - Percentage Composition Chemistry Answerkey and Bestseller Lists
5. Accessing Percentage Composition Chemistry Answerkey Free and Paid eBooks

- Percentage Composition Chemistry Answerkey Public Domain eBooks
- Percentage Composition Chemistry Answerkey eBook Subscription Services
- Percentage Composition Chemistry Answerkey Budget-Friendly Options
- 6. Navigating Percentage Composition Chemistry Answerkey eBook Formats
  - ePub, PDF, MOBI, and More
  - Percentage Composition Chemistry Answerkey Compatibility with Devices
  - Percentage Composition Chemistry Answerkey Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Percentage Composition Chemistry Answerkey
  - Highlighting and Note-Taking Percentage Composition Chemistry Answerkey
  - Interactive Elements Percentage Composition Chemistry Answerkey
- 8. Staying Engaged with Percentage Composition Chemistry Answerkey
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Percentage Composition Chemistry Answerkey
- 9. Balancing eBooks and Physical Books Percentage Composition Chemistry Answerkey
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Percentage Composition Chemistry Answerkey
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Percentage Composition Chemistry Answerkey
  - Setting Reading Goals Percentage Composition Chemistry Answerkey
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Percentage Composition Chemistry Answerkey
  - Fact-Checking eBook Content of Percentage Composition Chemistry Answerkey
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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