Name: Key

Date:

Honors Chemistry

Percent Yield Practice

 Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?

2 Al + 3 CuSO₄ - 3 Cu + Al₂(SO₄);

1. Start . 1 met Al . 3 met CuSO₄ 159.62 g (uSO₄ = 16.6 g CuSO₄.

26.98 g Al . 2 met Al . 1 met CuSO₄ = 16.6 g CuSO₄.

9.65 g CuSO₄ . 1 not CusO₄ . 3 met Cu . 63.55 g Cu = 3.84 g Cu

159.62 g CuSO₄ . 3 met Cu . 63.55 g Cu = 3.84 g Cu

159.62 g CuSO₄ . 3 met Cu . 1 met CuSO₄ . 1 met CuSO₄

3. Fyou start with 5.0 g of C₃H₈ and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?

ZNaCI + _ CaO - _ CaCI2 + Na2O

ZOg CaO . | mot CaO . / not tac 61.98 g Maz 0 = 200 g Maz 0 |

56.08 g CaO . | mot CaO . | mot

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Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, Prentice-Hall Staff

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