

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Percent yield is the amount of product that is actually produced compared to the theoretical yield. **Remember:** Limiting reagent is the reactant that will completely react and leave another reactant leftover.

Remember: Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Percent yield is the amount of product that is actually produced compared to the theoretical yield.



The reaction is balanced. The coefficients are 2, 1, and 2.

2H ₂	2 x 2.016 g/mol
O ₂	2 x 31.998 g/mol
2H ₂ O	2 x 18.015 g/mol
2H ₂	2 x 2.016 g/mol

1. 2.016 g/mol of hydrogen reacts with 31.998 g/mol of oxygen to produce 36.033 g/mol of water. **Remember:** Limiting reagent is the reactant that will completely react and leave another reactant leftover.



2. 36.033 g/mol of water is produced from 2.016 g/mol of hydrogen and 31.998 g/mol of oxygen.

3. 36.033 g/mol of water is produced from 2.016 g/mol of hydrogen and 31.998 g/mol of oxygen.



4. 36.033 g/mol of water is produced from 2.016 g/mol of hydrogen and 31.998 g/mol of oxygen. **Remember:** Limiting reagent is the reactant that will completely react and leave another reactant leftover.



Percent Yield And Limiting Reactants Quiz

Shasha Hu



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