

Percent Composition Practice 1.8 -- Answer Key

Objective: - Calculate the percent composition of a compound.

Directions: Determine the percent composition of each of the compounds below. (Show all work!)

1. HCl

$$\text{H: } 1.00\text{g} \times 1 = 1.01\text{g} \quad \left(\frac{1.01\text{g}}{35.46\text{g}} \right) \times 100\% = \boxed{2.73\%}$$

$$\text{Cl: } 35.45\text{g} \times 1 = 35.45\text{g} \quad \left(\frac{35.45\text{g}}{35.46\text{g}} \right) \times 100\% = \boxed{97.27\%}$$

36.46g

2. KMnO_4

$$\text{K: } 39.10\text{g} \times 1 = 39.10\text{g} \quad \left(\frac{39.10\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{24.74\%}$$

$$\text{Mn: } 54.94\text{g} \times 1 = 54.94\text{g} \quad \left(\frac{54.94\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{34.76\%}$$

$$\text{O: } 16.00\text{g} \times 4 = 64.00\text{g} \quad \left(\frac{64.00\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{40.50\%}$$

158.04g

3. $\text{Mg(NO}_3)_2$

$$\text{Mg: } 24.31\text{g} \times 1 = 24.31\text{g} \quad \left(\frac{24.31\text{g}}{148.31\text{g}} \right) \times 100\% = \boxed{16.39\%}$$

$$\text{N: } 14.01\text{g} \times 2 = 28.02\text{g} \quad \left(\frac{28.02\text{g}}{148.31\text{g}} \right) \times 100\% = \boxed{18.90\%}$$

$$\text{O: } 16.00\text{g} \times 6 = 96.00\text{g} \quad \left(\frac{96.00\text{g}}{148.31\text{g}} \right) \times 100\% = \boxed{64.72\%}$$

148.31g

4. $(\text{NH}_4)_3\text{PO}_4$

$$\text{N: } 14.01\text{g} \times 3 = 42.03\text{g} \quad \left(\frac{42.03\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{28.19\%}$$

$$\text{H: } 1.01\text{g} \times 12 = 12.12\text{g} \quad \left(\frac{12.12\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{8.13\%}$$

$$\text{P: } 30.97\text{g} \times 1 = 30.97\text{g} \quad \left(\frac{30.97\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{20.77\%}$$

$$\text{O: } 16.00\text{g} \times 4 = 64.00\text{g} \quad \left(\frac{64.00\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{42.92\%}$$

149.12g

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Table of Contents Percent Composition Answer Sheet

1. Understanding the eBook Percent Composition Answer Sheet
 - The Rise of Digital Reading Percent Composition Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Percent Composition Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Percent Composition Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Composition Answer Sheet
 - Personalized Recommendations
 - Percent Composition Answer Sheet User Reviews and Ratings
 - Percent Composition Answer Sheet and Bestseller Lists
5. Accessing Percent Composition Answer Sheet Free and Paid eBooks
 - Percent Composition Answer Sheet Public Domain eBooks
 - Percent Composition Answer Sheet eBook Subscription Services
 - Percent Composition Answer Sheet Budget-Friendly Options

6. Navigating Percent Composition Answer Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Percent Composition Answer Sheet Compatibility with Devices
 - Percent Composition Answer Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Percent Composition Answer Sheet
 - Highlighting and Note-Taking Percent Composition Answer Sheet
 - Interactive Elements Percent Composition Answer Sheet
8. Staying Engaged with Percent Composition Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Percent Composition Answer Sheet
9. Balancing eBooks and Physical Books Percent Composition Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Percent Composition Answer Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Percent Composition Answer Sheet
 - Setting Reading Goals Percent Composition Answer Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Percent Composition Answer Sheet
 - Fact-Checking eBook Content of Percent Composition Answer Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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