

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding



Bonding



There are two types of bonding

Definition of Polarity



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two multiple bonds** consist of one sigma and two pi bonds.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar bonds** are formed by the difference in electronegativity values of the two atoms.

• **Nonpolar bonds** are formed by the difference in electronegativity values of the two atoms, forming a nonpolar covalent bond. The important examples are the C-C and C-H bonds.

• **If the difference in electronegativity is between 0.5 and 1.7**, we have a polar covalent bond. The important examples are bonds for carbon to fluorine, N, O, F and other halogens.

• **A difference of 1.7 or higher** can be high that the electrons are not longer shared and a ionic bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



Electronegative atoms are color-coded in red, and electropositive atoms are color-coded in blue.

• The dipole moment is shown as a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• **δ^- (delta minus)** indicates an electron-deficient atom.

• **δ^+ (delta plus)** indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Cancel the vector as the other pointing point in the same way.



Non-polar molecule with polar bonds.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

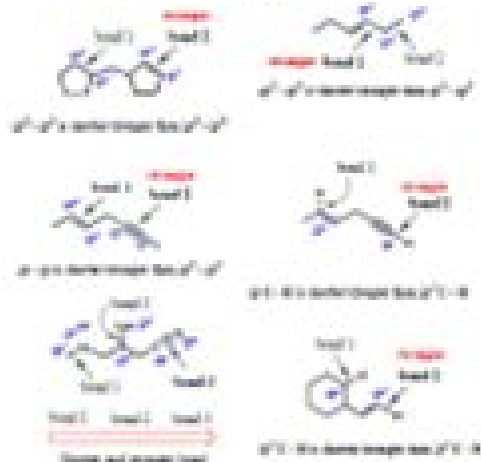


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	472 kJ/mol	505 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	410 kJ/mol	472 kJ/mol	505 kJ/mol

Higher s character = shorter and stronger bond

Examples



gaining information from the structure



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